

## Section 8.5

### Objectives

- **Describe** how electronegativity is used to determine bond type.
- **Compare and contrast** polar and nonpolar covalent bonds and polar and nonpolar molecules.
- **Generalize** about the characteristics of covalently bonded compounds.

## Review Vocabulary

**electronegativity:** the relative ability of an atom to attract electrons in a chemical bond

### New Vocabulary

polar covalent bond

## Electronegativity and Polarity

**MAIN Idea** A chemical bond's character is related to each atom's attraction for the electrons in the bond.

**Real-World Reading Link** The stronger you are, the more easily you can do pull-ups. Just as people have different abilities for doing pull-ups, atoms in chemical bonds have different abilities to attract (pull) electrons.

## Electron Affinity, Electronegativity, and Bond Character

The type of bond formed during a reaction is related to each atom's attraction for electrons. Electron affinity is a measure of the tendency of an atom to accept an electron. Excluding noble gases, electron affinity increases with increasing atomic number within a period and decreases with increasing atomic number within a group. The scale of electronegativities—shown in **Figure 8.20**—allows chemists to evaluate the electron affinity of specific atoms in a compound. Recall from Chapter 6 that electronegativity indicates the relative ability of an atom to attract electrons in a chemical bond. Note that electronegativity values were assigned, whereas electron affinity values were measured.

**Electronegativity** The version of the periodic table of the elements shown in **Figure 8.20** lists electronegativity values. Note that fluorine has the greatest electronegativity value (3.98), while francium has the least (0.7). Because noble gases do not generally form compounds, individual electronegativity values for helium, neon, and argon are not listed. However, larger noble gases, such as xenon, sometimes bond with highly electronegative atoms, such as fluorine.

**Figure 8.20** Electronegativity values are derived by comparing an atom's attraction for shared electrons to that of a fluorine's atom attraction for shared electrons. Note that the electronegativity values for the lanthanide and actinide series, which are not shown, range from 1.12 to 1.7.

1 <b>H</b> 2.20																	5 <b>B</b> 2.04	6 <b>C</b> 2.55	7 <b>N</b> 3.04	8 <b>O</b> 3.44	9 <b>F</b> 3.98	
3 <b>Li</b> 0.98	4 <b>Be</b> 1.57																	13 <b>Al</b> 1.61	14 <b>Si</b> 1.90	15 <b>P</b> 2.19	16 <b>S</b> 2.58	17 <b>Cl</b> 3.16
9 <b>K</b> 0.85	10 <b>Ca</b> 1.00	21 <b>Sc</b> 1.36	22 <b>Ti</b> 1.54	23 <b>V</b> 1.63	24 <b>Cr</b> 1.66	25 <b>Mn</b> 1.55	26 <b>Fe</b> 1.83	27 <b>Co</b> 1.88	28 <b>Ni</b> 1.91	29 <b>Cu</b> 1.90	30 <b>Zn</b> 1.65	31 <b>Ga</b> 1.87	32 <b>Ge</b> 2.05	33 <b>As</b> 2.18	34 <b>Se</b> 2.55	35 <b>Br</b> 2.96						
19 <b>Rb</b> 0.82	38 <b>Sr</b> 0.95	39 <b>Y</b> 1.22	40 <b>Zr</b> 1.33	41 <b>Nb</b> 1.6	42 <b>Mo</b> 2.16	43 <b>Tc</b> 2.09	44 <b>Ru</b> 2.2	45 <b>Rh</b> 2.28	46 <b>Pd</b> 2.20	47 <b>Ag</b> 1.93	48 <b>Cd</b> 1.69	49 <b>In</b> 1.78	50 <b>Sn</b> 1.96	51 <b>Sb</b> 2.05	52 <b>Te</b> 2.1	53 <b>I</b> 2.66						
37 <b>Ba</b> 0.82	56 <b>La</b> 1.10	57 <b>Ce</b> 1.1	72 <b>Hf</b> 1.3	73 <b>Ta</b> 1.5	74 <b>W</b> 1.7	75 <b>Re</b> 1.9	76 <b>Os</b> 2.2	77 <b>Ir</b> 2.2	78 <b>Pt</b> 2.2	79 <b>Au</b> 2.4	80 <b>Hg</b> 1.9	81 <b>Tl</b> 1.8	82 <b>Pb</b> 1.8	83 <b>Bi</b> 1.9	84 <b>Po</b> 2.0	85 <b>At</b> 2.2						
55 <b>Cs</b> 0.79	88 <b>Ra</b> 0.9	89 <b>Ac</b> 1.1																				

# Electronegativity And Polarity Study Guide

**Leo J. Malone, Theodore O. Dolter**



## Electronegativity And Polarity Study Guide:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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A Self-study Guide to the Principles of Organic Chemistry Jiben Roy, 2013 A Self-Study Guide to the Principles of Organic Chemistry: Key Concepts, Reaction Mechanisms, and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily, as well as build a strong foundation for future study. Starting with the definition of an atom, the author explains molecules, electronic configuration, bonding, hydrocarbons, polar reaction mechanisms, stereochemistry, reaction varieties, organic spectroscopy, aromaticity, and aromatic reactions, biomolecules, organic polymers, and a synthetic approach to organic compounds. The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood. Each chapter ends with a list of frequently asked questions and answers, followed by additional practice problems. Answers are included in the Appendix.

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