



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2O ionic b. CH_4 covalent c. HCl covalent
d. KNO_3 ionic e. H_2O covalent

2. Write the formula for the following molecular compounds.

- a. sulfur dioxide b. carbon monoxide c. dinitrogen
d. sulfur trioxide e. carbon dioxide f. nitrogen

3. Write the molecular formula for each of the following compounds. Name each.

4. Write the molecular formula for each of the following compounds. Name each. Write the molecular formula for each of the following compounds. Name each.

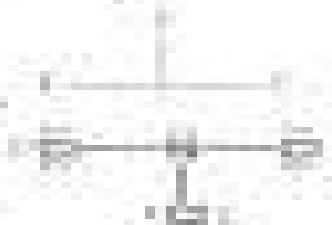
5. Write the molecular formula for each of the following compounds. Name each. Write the molecular formula for each of the following compounds. Name each.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .



4. Draw the Lewis structure for acetylene, C_2H_2 .

5. Draw the Lewis structure for each of the following compounds. Name each.



6. Draw the Lewis structure for each of the following compounds. Name each.

SECTION 8.3 BONDING: THE ORBITAL

1. Write the Lewis structure for each of the following compounds. Name each. Write the Lewis structure for each of the following compounds. Name each.

2. Write the Lewis structure for each of the following compounds. Name each. Write the Lewis structure for each of the following compounds. Name each.

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SJ Ball



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