

COMMON CHEMISTRY TERMS

Matter

Atom - basic unit of an element
Atom - negatively charged ion
Compound - group of two or more different atoms bonded together (H_2O , CO_2 , CH_4)
Cation - positively charged ion
Chemistry - the study of matter
Diatomic - two highly bonded atoms in pairs (= diatomic)
Electron - tiny negative electrically outside the nucleus of an atom
Ions - charged particles due to loss or gain of electrons
Isomers - multiple forms of same chemical with different molecular structures
Matter - anything that takes up space and has mass
Molecule - group of atoms bonded together (H_2 , O_2 , H_2O , CO_2)
Nucleus - central of an atom consisting of protons and neutrons
Neutron - neutral particle in the nucleus of an atom
Phase - a form of matter solid, liquid, gas, or plasma
Protons - positively charged inside nucleus

Chemical Reaction Vocabulary

Catalyst - something that speeds up the rate of a chemical reaction
Endothermic - a reaction that absorbs heat
Entropy - degree of disorder in a system
Enzyme - a biological catalyst
Equilibrium - when the rate of the forward reaction equals the rate of the reverse reaction
Exothermic - a reaction that releases heat
Reactants - substances that undergo a change in a reaction (on the left of the arrow in a chemical equation)
Preproduct - a solid that forms from a solution
Products - substance that is formed in a chemical reaction (on the right of the arrow in a chemical equation)

Common Types of Reactions

Combustion: $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ - exothermic
Decomposition: $\text{AB} \rightarrow \text{A} + \text{B}$
Dehydration - a reaction that results in loss of water (water appears in the products)
Double Displacement: $\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$
Hydration - a reaction where a chemical substance reacts with water to form a product (water is in the reactants)
Neutralization: $\text{Acid} + \text{Base} \rightarrow \text{H}_2\text{O} + \text{salt}$
Synthesis: $\text{A} + \text{B} \rightarrow \text{AB}$
Single Displacement: $\text{A} + \text{BC} \rightarrow \text{AC} + \text{B}$

Solutions, Acids, and Bases

Acid - molecule that can either donate a proton (H^+) or form a constant bond by accepting a pair of electrons. Types include:
dicarboxylic acids - donate a proton
acids - form constant bond by accepting electron pair
Amphoteric - substance that can react as an acid and a base
Base - molecule that can increase concentration of hydroxide (OH^-) or accept a proton (H^+) or donate a pair of electrons. Types include:
alkaline bases - increase concentration of hydroxide ions
dicarboxylic bases - accept a proton
acids - donate an electron pair
Concentration - how much of a substance is mixed with another substance
Molarity - the number of moles of a substance in a certain volume of solution. An example of a concentration
Mole - standard unit used to measure the amount of any substance (1 mole = 6.022 x 10²³ individual components)
pH - a measure of the acidity of a solution
Solute - what is dissolved in the solvent
Solvent - homogeneous mixture with 7 or more components
Solvent - what a solute is dissolved in

Periodic Trends

Atomic radius - size of an atom
Electron Affinity - amount of energy released when an atom accepts an electron
Electronegativity - ability of an atom to attract shared electrons in bond
Ionization Energy - amount of energy required to remove an electron from an atom

Bonding + Intermolecular Forces

Covalent Bond - a bond where electrons are shared between atoms
Dipole-dipole forces - attractive intermolecular forces between positive end of a polar molecule and the negative end of another polar molecule
Hydrogen bonding - a type of dipole-dipole intermolecular force that occurs with the following bond types: H-F, H-O, and H-N
Intermolecular force - the forces that exist between separate atoms and molecules
Ionic Bond - bond formed by the transfer of electrons between atoms
Lewis Structure - simplified representation of valence electrons for an atom
Lewis structure - weakest intermolecular force between atoms and molecules
Valence electrons - electrons in outermost shell of an atom

Chemistry Vocabulary Study Guide

F Rizvi



Chemistry Vocabulary Study Guide:

Chemistry Vocabulary Study Guide Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Chemistry Vocabulary Study Guide**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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