- 9. $Fe_2O_3 \rightarrow Fe_1e_2O_2$
- 10. $P + O_2 \rightarrow P_2O_5$
- 11. $\underline{\hspace{0.1cm}}$ Mg + $\underline{\hspace{0.1cm}}$ HCl \rightarrow $\underline{\hspace{0.1cm}}$ MgCl₂ + $\underline{\hspace{0.1cm}}$ H₂
- 12. $\underline{\hspace{1cm}}$ H₂ + $\underline{\hspace{1cm}}$ N₂ \rightarrow $\underline{\hspace{1cm}}$ NH₃
- 13. $\underline{\hspace{0.2cm}}$ BaCl₂ + $\underline{\hspace{0.2cm}}$ H₂SO₄ \rightarrow $\underline{\hspace{0.2cm}}$ BaSO₄ + $\underline{\hspace{0.2cm}}$ HCl
- - b) Find the molar mass of these reactants.
 - c) How many moles of ZnCl2 would be in 25 g?
 - d) How much mass would 0.55 moles of (NH₄)₂S have?

Chemistry Unit 7 Reaction Equations

Padhraic Smyth

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