

## 15 Hydrocarbons

### Alkanes

#### Chemical Reactivity

Alkanes do not react with many reagents.

This is because the C-C bond and the C-H bond are relatively strong and so cannot form stronger bonds with other elements that easily and there are no polar side groups.

Their main reactions are combustion reactions in air although they can react with halogens in the presence of UV light.

#### Combustion

##### Complete Combustion

In excess oxygen alkanes will burn with complete combustion.

The products of complete combustion are  $\text{CO}_2$  and  $\text{H}_2\text{O}$ .



Alkanes readily burn in the presence of oxygen. This combustion of alkanes is highly exothermic, explaining their use as **fuels**.

#### Reaction of alkanes with bromine / chlorine in UV light

##### Synthesis of chloroalkanes

In the presence of **UV light** alkanes react with chlorine to form a mixture of products with the halogens substituting hydrogen atoms.

##### Overall Reaction



*This is the overall reaction, but a more complex mixture of products is actually formed.*

*To understand this reaction fully we must look in detail at how it proceeds step by step. This is called its **mechanism**.*

The **MECHANISM** for this reaction is called a **FREE RADICAL SUBSTITUTION**.

It proceeds via a series of steps:  
**STEP ONE: Initiation**  
**STEP TWO: Propagation**  
**STEP THREE: Termination**

##### STEP ONE Initiation

Essential condition: UV light



The UV light supplies the energy to break the Cl-Cl bond. It is broken in preference to the others as it is the weakest.

The bond has broken in a process called **homolytic fission**.

*each atom gets **one** electron from the covalent bond*

When a bond breaks by homolytic fission it forms **Free Radicals**.

**Free Radicals** do not have a charge and are represented by a



##### DEFINITION

A **Free Radical** is a reactive species which possess an **unpaired electron**.

##### STEP TWO Propagation



The chlorine free radicals are very reactive and remove an H from the methane leaving a methyl free radical

The methyl free radical reacts with a  $\text{Cl}_2$  molecule to produce the main product and another Cl free radical

All propagation steps have a **free radical** in the **reactants** and in the **products**.

As the Cl free radical is regenerated, it can react with several more alkane molecules in a **CHAIN REACTION**.

# Chemistry Study Guide Hydrocarbons

**S Ben Porath**



## **Chemistry Study Guide Hydrocarbons:**

The book delves into Chemistry Study Guide Hydrocarbons. Chemistry Study Guide Hydrocarbons is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemistry Study Guide Hydrocarbons, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
  - Chapter 1: Introduction to Chemistry Study Guide Hydrocarbons
  - Chapter 2: Essential Elements of Chemistry Study Guide Hydrocarbons
  - Chapter 3: Chemistry Study Guide Hydrocarbons in Everyday Life
  - Chapter 4: Chemistry Study Guide Hydrocarbons in Specific Contexts
  - Chapter 5: Conclusion
2. In chapter 1, this book will provide an overview of Chemistry Study Guide Hydrocarbons. The first chapter will explore what Chemistry Study Guide Hydrocarbons is, why Chemistry Study Guide Hydrocarbons is vital, and how to effectively learn about Chemistry Study Guide Hydrocarbons.
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4. In chapter 3, the author will examine the practical applications of Chemistry Study Guide Hydrocarbons in daily life. The third chapter will showcase real-world examples of how Chemistry Study Guide Hydrocarbons can be effectively utilized in everyday scenarios.
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6. In chapter 5, the author will draw a conclusion about Chemistry Study Guide Hydrocarbons. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Chemistry Study Guide Hydrocarbons.

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