

## Unit 5 – Electrochemistry Test

NAME: \_\_\_\_\_

K/U	C	A	T/I
/8	/10	/7	/4

### Part A: Knowledge & Understanding – Multiple Choice (8 Marks)

*Identify the choice that best completes the statement or answers the question.*

- Which of the following substances would make a poor electrolyte solution?  
a) Lead (II) Nitrate                      c) Silver Bromide  
b) Potassium Chloride                d) Ammonium Phosphate
- Which of the following solid metals could act as the anode if paired with cadmium ( $\text{Cd(s)}/\text{Cd}^{2+}$ ) in a galvanic cell?  
a) Lead                                      c) Zinc  
b) Copper                                  d) Tin
- The proper cell notation for a galvanic cell composed of a chromium electrode and a cadmium electrode would be:  
a)  $\text{Cr(s)} \mid \text{Cr}^{3+}(\text{aq}) \parallel \text{Cd}^{2+}(\text{aq}) \mid \text{Cd(s)}$   
b)  $\text{Cd(s)} \mid \text{Cd}^{2+}(\text{aq}) \parallel \text{Cr}^{3+}(\text{aq}) \mid \text{Cr(s)}$   
c)  $\text{Cr}^{3+}(\text{aq}) \mid \text{Cr(s)} \parallel \text{Cd(s)} \mid \text{Cd}^{2+}(\text{aq})$   
d)  $\text{Cd}^{2+}(\text{aq}) \mid \text{Cd(s)} \parallel \text{Cr(s)} \mid \text{Cr}^{3+}(\text{aq})$
- Which of the following metals is the least susceptible to rusting?  
a) Nickel  
b) Copper  
c) Gold  
d) Zinc
- Which of the following is an example of an alloy?  
a) Bromine  
b) Iron  
c) Bronze  
d) Copper Nitrate

# Electrochemistry Test

**R Sanford**



## Electrochemistry Test:

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