

Counting by Measuring Mass

Purpose

To determine the mass of several samples of chemical compounds and use the data to count atoms

Materials

H₂O, NaCl, and CaCO₃
plastic spoon
weighing paper
balance

Safety



Procedure

Measure the mass of one level teaspoon of water (H₂O), sodium chloride (NaCl), and calcium carbonate (CaCO₃). Make a table similar to the one below.

	H ₂ O(l)	CaCl ₂	Na ₂ CO ₃
Mass (g)	2.42 g	3.95 g	2.30 g
Molar mass (g/mol)	18.0 g/mol	111.1 g/mol	106.0 g/mol
Moles of each compound	0.146 mol	0.0354 mol	0.0217 mol
Moles of each element	0.292 mol H 0.146 mol O	0.0354 mol Ca 0.0708 mol Cl	0.0434 mol Na 0.0217 mol C 0.0651 mol O
Atoms of each element	1.76 × 10 ²³ H 1.79 × 10 ²³ O	2.15 × 10 ²³ Ca 4.26 × 10 ²³ Cl	2.61 × 10 ²³ Na 1.31 × 10 ²³ C 3.92 × 10 ²³ O



$$2 \times 23.0 + 12.0 + 3 \times 16.0 = 106.0$$

$$2.30 \text{ g} \left(\frac{1 \text{ mol}}{106.0 \text{ g}} \right) = 0.0217$$

$$2.30 \text{ g} \left(\frac{1 \text{ mol}}{106.0 \text{ g}} \right)$$

1. **Calculate** Determine the number of moles of H₂O contained in one level teaspoon.

$$\text{moles of H}_2\text{O} = \text{g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}}$$

Repeat for the remaining compounds. Use the periodic table to calculate the molar masses of NaCl and CaCO₃.

2. **Calculate** Determine the number of moles of each element present in the teaspoon-sized sample of H₂O.

$$\text{moles of H} = \text{mol H}_2\text{O} \times \frac{2 \text{ mol H}}{1 \text{ mol H}_2\text{O}}$$

Repeat for the other compounds in your table.

3. **Calculate** Determine the number of atoms of each element present in the teaspoon-sized sample of H₂O.

$$\text{atoms of H} = \text{mol H} \times \frac{6.02 \times 10^{23} \text{ atoms H}}{1 \text{ mol H}}$$

Repeat for the other compounds in your table.

Counting By Measuring Mass Lab Answer Key

Karen Timberlake



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Table of Contents Counting By Measuring Mass Lab Answer Key

1. Understanding the eBook Counting By Measuring Mass Lab Answer Key
 - The Rise of Digital Reading Counting By Measuring Mass Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Counting By Measuring Mass Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Counting By Measuring Mass Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Counting By Measuring Mass Lab Answer Key
 - Personalized Recommendations
 - Counting By Measuring Mass Lab Answer Key User Reviews and Ratings
 - Counting By Measuring Mass Lab Answer Key and Bestseller Lists
5. Accessing Counting By Measuring Mass Lab Answer Key Free and Paid eBooks
 - Counting By Measuring Mass Lab Answer Key Public Domain eBooks
 - Counting By Measuring Mass Lab Answer Key eBook Subscription Services
 - Counting By Measuring Mass Lab Answer Key Budget-Friendly Options

6. Navigating Counting By Measuring Mass Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Counting By Measuring Mass Lab Answer Key Compatibility with Devices
 - Counting By Measuring Mass Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Counting By Measuring Mass Lab Answer Key
 - Highlighting and Note-Taking Counting By Measuring Mass Lab Answer Key
 - Interactive Elements Counting By Measuring Mass Lab Answer Key
8. Staying Engaged with Counting By Measuring Mass Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Counting By Measuring Mass Lab Answer Key
9. Balancing eBooks and Physical Books Counting By Measuring Mass Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Counting By Measuring Mass Lab Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Counting By Measuring Mass Lab Answer Key
 - Setting Reading Goals Counting By Measuring Mass Lab Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Counting By Measuring Mass Lab Answer Key
 - Fact-Checking eBook Content of Counting By Measuring Mass Lab Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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