

# Electrode Potential

- Definition:
  - the difference in potential between the metal/metal ion system and another system
- Measured in volts
- System used for comparison: standard hydrogen electrode
- Electrode potential refers to the reduction reaction (electrons appear on the left-hand side)



\* The more positive (or the less negative) the electrode potential, the easier it is to reduce the ions on the left



equilibrium shift to right, reduction is favorable



equilibrium shift to left, reduction is not favorable

POSITIVE MEANS EASY TO REDUCE, NEGATIVE MEANS DIFFICULT TO REDUCE

Practice :-

10 (c) If positive  $E^\circ$  means easy to reduce  $\rightarrow$  good O.A.

Best reducing agent  $\rightarrow$  most  $-E^\circ$  : Zn

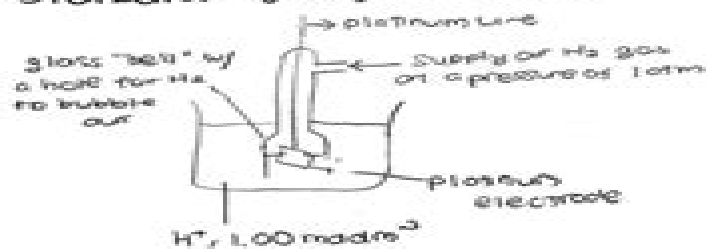
(b)  $Zn^{2+}$

(c) most reactive: most easy to oxidise

Zn

(d)  $Ag^+$

## Standard Hydrogen Electrode



This electrode consists of:

- $H_2(g)$  at 101 kPa pressure - in equilibrium w/
- $H^+$  ions of conc.  $1.00 \text{ mol dm}^{-3}$
- a platinum electrode covered w/ platinum black in contact w/ the  $H_2$  gas and the  $H^+$  ions

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