

9. $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g})$ $.10$ $.02$ @ 25°C
 At 25°C , 0.11 mole of N_2O_4 reacts to form 0.10 mol of N_2O_4 and 0.02 mole of NO_2 . At 90°C , 0.11 mole of N_2O_4 forms 0.050 mole of N_2O_4 and 0.12 mole of NO_2 . From these data we can conclude $.05$ $.12$ @ 90°C
 a) N_2O_4 molecules react by a second order rate law.
 b) N_2O_4 molecules react by a first order rate law.
 c) the reaction is ~~exothermic~~ ^{endo}.
 d) N_2O_4 molecules react faster at 25°C than at 90°C .
 e) the equilibrium constant for the reaction above increases with an increase in temperature. *more products at 90°C than 25°C*

10. For the equilibrium system
 $\text{H}_2\text{O}(\text{g}) + \text{CO}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{CO}_2(\text{g})$ $+\text{Heat}$
 $\Delta H = -42 \text{ kJ/mol}$
 K_c equals 0.62 at 1260 K . If 0.10 mole each of H_2O , CO , H_2 and CO_2 (each at 1260 K) were placed in a 1.0-Liter flask at 1260 K , when the system came to equilibrium...

	The temperature would	The mass of CO would
a)	decrease	increase
b)	decrease	decrease
c)	remain constant	increase
d)	increase	decrease
e)	increase	increase

$$Q = \frac{(.10)(.10)}{(.10)(.10)} = 1 \quad K_c = .62$$

$Q > K_c$ will shift to
 To use up products
 and make reactants
 Temp $\downarrow$ CO $\uparrow$

11. For the reaction system,
 $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g}) + \text{heat}$
 the conditions that would favor maximum conversion of the reactants to products would be *Fewer moles of gas on product side*
 a) high-temperature and high pressure
 b) high-temperature, pressure unimportant
 c) high-temperature and low pressure
 d) low temperature and high pressure
 e) low temperature and low pressure
12. Solid HgO , liquid Hg , and gaseous O_2 are placed in a glass bulb and are allowed to reach equilibrium at a given temperature.
 $2\text{HgO}(\text{s}) \rightleftharpoons 2\text{Hg}(\text{l}) + \text{O}_2(\text{g})$ $\Delta H = +43.4 \text{ kcal}$
 The mass of HgO in the bulb could be increased by $\leftarrow$ shift left
 a) adding more Hg . *NO CHANGE*
 b) removing some O_2 . $\rightarrow$
 c) reducing the volume of the bulb. $\rightarrow$ *Fewer moles of gas on reactant side*
 d) increasing the temperature. $\rightarrow$
 e) removing some Hg . *NO CHANGE*

Answers: (Please use CAPITAL letters)

1.	E
2.	D
3.	C
4.	B
5.	D
6.	D

7.	D
8.	A
9.	E
10.	A
11.	D
12.	C

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