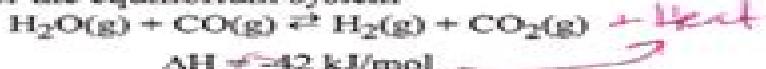


9.



- At 25°C , 0.11 mole of N_2O_4 reacts to form 0.10 mol of N_2O_4 and 0.02 mole of NO_2 . At 90°C , 0.11 mole of N_2O_4 forms 0.050 mole of N_2O_4 and 0.12 mole of NO_2 . From these data we can conclude
- a) N_2O_4 molecules react by a second order rate law.
 - b) N_2O_4 molecules react by a first order rate law.
 - c) the reaction is *endo* exothermic.
 - d) N_2O_4 molecules react faster at 25°C than at 90°C .
 - e) the equilibrium constant for the reaction above increases with an increase in temperature. *more products at 90°C than 25°C*

10. For the equilibrium system



K_c equals 0.62 at 1260 K . If 0.10 mole each of H_2O , CO , H_2 and CO_2 (each at 1260 K) were placed in a 1.0-Liter flask at 1260 K , when the system came to equilibrium...

| The temperature would | The mass of CO would |
|-----------------------|----------------------|
| decrease | increase |
| decrease | decrease |
| remain constant | increase |
| increase | decrease |
| increase | increase |

$$Q = \frac{(0.10)(0.10)}{(0.10)(0.10)} = 1 \quad K_c = 0.62$$

$Q > K_c$ will start to \leftarrow
To use up products and make reactants
Temp \downarrow CO \uparrow

11. For the reaction system,



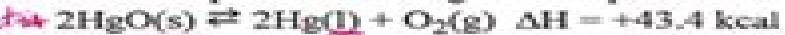
the conditions that would favor maximum conversion of the reactants to products would be

fewer moles of gas on product side

- a) high-temperature and high pressure
- b) high-temperature, pressure unimportant
- c) high-temperature and low pressure
- d) low temperature and high pressure
- e) low temperature and low pressure

12. Solid HgO , liquid Hg , and gaseous O_2 are

placed in a glass bulb and are allowed to reach equilibrium at a given temperature.



The mass of HgO in the bulb could be increased by

shut left

a) adding more Hg . *NO CHANGE*

b) removing some O_2 . *NO CHANGE*

c) reducing the volume of the bulb. *NO CHANGE*

d) increasing the temperature. *NO CHANGE*

e) removing some Hg . *NO CHANGE*

Answers: (Please use **CAPITAL** letters)

| | |
|----|---|
| 1. | E |
| 2. | D |
| 3. | C |
| 4. | B |
| 5. | D |
| 6. | D |

| | |
|-----|---|
| 7. | D |
| 8. | A |
| 9. | E |
| 10. | A |
| 11. | D |
| 12. | C |

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