hmPhysical Properties of Gases

- Expansion Gasses expand to fill entirety of a container
- Fluidity Gases flow past one another
- Low Density Gases are less dense than liquid and solids
- Compressibility Gases can be compressed down into small spaces
- Diffusion Gas molecules disperse through a container
- Effusion Gas passes through smaller openings than the free path of the particles

Kinetic Molecular Theory

- All particles of matter in constant motion
- Gas molecules are tiny and far apart
 - Low Density and compressibility
- Collide with each other in elastic collisions
 - Pressure
- Do not attract or repel each other
 - Fluidity and expansion
- Constant motion
 - Diffusion, effusion, pressure
- Kinetic energy directly proportional to temperature

4 main gas variables

- Pressure
- Temperature
- Volume
- Moles

Gas Laws

- B's law: P1"V1=P2"V2
 - P and V are inversely proportional
- C's Law: V1/T1= V2/T2
 - V and T are directly proportional
- L's Law: P1/T1=P2/T2
 - P and T are directly proportional
- Combined gas law
 - P1*V1/T1=P2*V2/T2
- Ideal Gas Laws
 - PV=nRT
 - R= 0.0821 atm
 - R=8.31 kPa
 - R= 62.4 mmHg
 - kPa, Liters, K.
- Partial Pressure Law
 - P total = P1 + P2+ P3
 - Ptotal = Pressure of gas + pressure of water vapor
- Partial Pressure Law w/ Mass and Density
 - PM=DRT
 - Density

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