

# Chemistry Spring Final Study Guide

### Unit 10: Acids & Bases

Use the following information and complete the structure of Acids & Bases.

Acids	Bases
<p><b>Unit 10: Acids &amp; Bases</b></p> <p>Draw a line on the graph and indicate its dependence for each property: Temperature, Volume, and Pressure below.</p> <p>Label the y-axis.</p> <p>Draw the line for:</p> <ol style="list-style-type: none"> <li>Boiling point of the gas phase</li> <li>Boiling point of the liquid phase</li> <li>A mixture of the gas and liquid</li> <li>The gas in air</li> <li>The condensation</li> </ol>	<p><b>Unit 10: Acids &amp; Bases</b></p> <p>Draw the line for each property: Temperature, Volume, and Pressure below.</p> <p>Label the y-axis.</p> <p>Draw the line for:</p> <ol style="list-style-type: none"> <li>Boiling point of the gas phase</li> <li>Boiling point of the liquid phase</li> <li>A mixture of the gas and liquid</li> <li>The gas in air</li> <li>The condensation</li> </ol>

### Unit 11: Chemical Equilibrium

Draw the line for each property: Temperature, Volume, and Pressure below.

Label the y-axis.

Draw the line for:

- Boiling point of the gas phase
- Boiling point of the liquid phase
- A mixture of the gas and liquid
- The gas in air
- The condensation

### Unit 12: Chemical Equilibrium

Draw the line for each property: Temperature, Volume, and Pressure below.

Label the y-axis.

Draw the line for:

- Boiling point of the gas phase
- Boiling point of the liquid phase
- A mixture of the gas and liquid
- The gas in air
- The condensation

## Chemistry Spring Final Study Guide

Name: \_\_\_\_\_

### Unit 1: Measurement & Analysis

Mark the best answer from those provided to each question.

Mark the best answer from those provided to each question.

Determine the number of Significant Figures in each of the following numbers:

Mass = 1.23456789 g	Mass = 1.23456789 g
Volume = 1.23456789 L	Volume = 1.23456789 L

Convert the following into Scientific Notation:

Mass = 1.23456789 g	Mass = 1.23456789 g
Volume = 1.23456789 L	Volume = 1.23456789 L

Convert the following into Decimal Notation:

Mass = 1.23456789 g	Mass = 1.23456789 g
Volume = 1.23456789 L	Volume = 1.23456789 L

Calculate the molar mass of each compound:

Mass = 1.23456789 g	Mass = 1.23456789 g
Volume = 1.23456789 L	Volume = 1.23456789 L

# Chemistry Study Guide Spring 2013

**Jicheng Xie**



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