

Nuclear Stability

We can establish a pattern of reactions between protons and neutrons in nuclear reactions



The question is why, for example, some nuclei undergo an alpha decay and not a beta decay. This has to do with the **neutrons to protons ratio** (n/p) in the nuclei. When the nuclei of all 1000 known isotopes are plotted (based on their number of protons and neutrons), a **band (belt) of stability** is obtained (blue region). The nuclei in this region are stable. The isotopes that fall outside of this region (are off-balance) are radioactive and decay spontaneously. Example 11 is the last element on the band of stability, so all nuclei with 84 or more protons are radioactive. The type of radioactivity depends on the n/p

- The nuclei above the band of stability tend to decrease the n/p ratio.
- The nuclei below the band of stability tend to increase the n/p ratio.



Nuclei with the **magic numbers** of 2, 8, 20, 28, 50, or 82 protons or 2, 8, 20, 28, 50, 82, or 126 neutrons, and in general, with even numbers of protons, neutrons are more stable.

Rates of Radioactive Decay

Radioactive decay obeys first-order kinetics laws. And although the half-life of radioactive elements may vary enormously (10^2 second to billions of years), it is independent of the quantity of the element. This advantage is used in **radiocarbon dating** which is used for **estimating the ages of fossils and artifacts**. The technique is based on the rate of ${}^{14}\text{C}$ decay which spreads from the atmosphere to living organisms.



The half-life of carbon-14 decay is averaged by emission of a β^- particle is 5715 years.



The ${}^{14}\text{C} : {}^{12}\text{C}$ ratio is constant in living organisms. However, when a living organism dies, the ${}^{14}\text{C} : {}^{12}\text{C}$ ratio begins to decrease, and so does the decay rate of ${}^{14}\text{C}$. The relationship of the quantity or rates of decay is given by the first-order rate law:

$$\ln \left(\frac{N}{N_0} \right) = -kt \quad \ln \left(\frac{\text{rate}_t}{\text{rate}_0} \right) = -kt$$

Example: A sample of wood from an Egyptian mummy case gives a count of 9.4 counts/g, this is the current decay rate determined experimentally. How old is the wood if its initial decay rate is 15.3 counts/g?

Solution: We are going to use the integrated rate law of first-order reactions. The rates are given, so we need to determine the rate constant of the decay.

The half-life of first-order reactions is $t_{1/2} = 0.693/k$. Therefore, k is:

$$k = \frac{0.693}{t_{1/2}} = \frac{0.693}{5715 \text{ y}} = 1.210 \times 10^{-4} \text{ y}^{-1}$$

$$\ln \left(\frac{9.4}{15.3} \right) = (1.210 \times 10^{-4}) \times t \quad \Rightarrow \quad t = \frac{\ln \left(\frac{9.4}{15.3} \right)}{(1.210 \times 10^{-4})} = \boxed{4,829 \text{ years}}$$

Energy Changes in Nuclear Reactions

The energies involved with nuclear reactions are extremely large. For example, the energy change in this reaction is $-4.1 \times 10^{12} \text{ J/mol}$



For comparison, the energy change for combustion of 1 mol CH_4 is 890 kJ/mol.

Chemistry Study Guide Nuclear Chemistry

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CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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Fundamental Concepts of Applied Chemistry Jayashree Ghosh, 2010-12 During the past few decades the growth of applied chemistry has been phenomenal and its applications have an expansive field including Chemical and Medico Biological disciplines I take pleasure in presenting the book Fundamental concepts of applied chemistry The book is published to provide a concise text book that encompasses important branches like pharmaceutical Biological polymer leather and Agricultural Chemistry *Kaplan SAT Subject Test Chemistry 2015-2016* Kaplan Test Prep, 2015-03-03 Essential strategies practice and review to ace the SAT Subject Test Chemistry Getting into a top college has never been more difficult Students need to distinguish themselves from the crowd and scoring well on a SAT Subject Test gives students a competitive edge Kaplan s SAT Subject Test Chemistry is the most up to date guide on the market with complete coverage of both the content review and strategies students need for success on test day Kaplan s SAT Subject Test Chemistry features A full length diagnostic test Full length practice tests Focused chapter summaries highlights and quizzes Detailed answer explanations Proven score raising strategies End of chapter quizzes Kaplan is serious about raising students scores we guarantee students will get a higher score United States Armed Forces Institute Catalog United States Armed Forces Institute, 1963

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