

Answer the following questions:

1. State the following stoichiometric equations:
a) $\text{CH}_3\text{Cl} + \text{O}_2 \rightarrow \text{CH}_2\text{Cl} + \text{HCl}$ $\Delta H = -100 \text{ kJ}$
b) In more reactant \rightarrow faster reaction rate
 CH_3Cl reacts \rightarrow CH_2Cl and HCl are products.
2. State whether the overall stoichiometric reaction equation between:
a) CH_3Cl and O_2 \rightarrow CH_2Cl and HCl
b) CH_3Cl and the overall stoichiometric reaction between:
 $\text{Al}_2\text{O}_3 + \text{Cl}_2 \rightarrow \text{AlCl}_3$ and O_2
c) CH_3Cl and H_2O \rightarrow CH_2Cl and H_2
3. Consider the following stoichiometric equations: $\text{CH}_3\text{Cl} + \text{O}_2 \rightarrow \text{CH}_2\text{Cl} + \text{HCl}$ and $\text{CH}_3\text{Cl} + \text{H}_2\text{O} \rightarrow \text{CH}_2\text{Cl} + \text{H}_2$.
a) Which decomposition reaction has the higher activation energy?
b) Which reaction has the higher overall reaction rate?
c) CH_3Cl reacts \rightarrow CH_2Cl and HCl
4. If it is found that the reaction rates of decomposition reactions can be altered or altered by the reaction
time, then the $\text{Al}_2\text{O}_3 + \text{Cl}_2 \rightarrow \text{AlCl}_3$ reaction is first-order.
However, the $\text{CH}_3\text{Cl} + \text{H}_2\text{O} \rightarrow \text{CH}_2\text{Cl} + \text{H}_2$ reaction is second-order.
5. Below each reaction write the general rate law equation for the reaction.
a) $\text{CH}_3\text{Cl} + \text{H}_2\text{O} \rightarrow \text{CH}_2\text{Cl} + \text{H}_2$
b) $\text{CH}_3\text{Cl} + \text{O}_2 \rightarrow \text{CH}_2\text{Cl} + \text{HCl}$
6. Below, given a series of reaction conditions, write down the overall reaction rate expression
for each reaction:
a) $\text{CH}_3\text{Cl} + \text{H}_2\text{O} \rightarrow \text{CH}_2\text{Cl} + \text{H}_2$
i) CH_3Cl is present in excess.
ii) H_2O is present in excess.
iii) CH_3Cl reacts with H_2O at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
- b) $\text{CH}_3\text{Cl} + \text{O}_2 \rightarrow \text{CH}_2\text{Cl} + \text{HCl}$
i) CH_3Cl reacts with O_2 at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
ii) O_2 reacts with CH_3Cl at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
- c) $\text{CH}_3\text{Cl} + \text{H}_2\text{O} \rightarrow \text{CH}_2\text{Cl} + \text{H}_2$
i) CH_3Cl reacts with H_2O at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
ii) H_2O reacts with CH_3Cl at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
- d) $\text{CH}_3\text{Cl} + \text{O}_2 \rightarrow \text{CH}_2\text{Cl} + \text{HCl}$
i) CH_3Cl reacts with O_2 at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
ii) O_2 reacts with CH_3Cl at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
- e) $\text{CH}_3\text{Cl} + \text{H}_2\text{O} \rightarrow \text{CH}_2\text{Cl} + \text{H}_2$
i) CH_3Cl reacts with H_2O at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
ii) H_2O reacts with CH_3Cl at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
- f) $\text{CH}_3\text{Cl} + \text{O}_2 \rightarrow \text{CH}_2\text{Cl} + \text{HCl}$
i) CH_3Cl reacts with O_2 at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.
ii) O_2 reacts with CH_3Cl at a rate of $10 \text{ mol L}^{-1} \text{ s}^{-1}$.

Chemistry Thermochemical Equations Study Guide

Answers

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Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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A-Level Study Guide Chemistry Ed H2.2 CS Toh, 2016-03-08 This is an ebook version of the A Level Study Guide Chemistry Higher 2 Ed H2 2 published by Step by Step International Pte Ltd For the revised Higher 2 H2 syllabus with first exam in 2017 This ebook gives concise illustrated notes and worked examples It is intended as a study guide for readers who have studied the O Level Chemistry or the equivalent It contains material that most readers should want to take note of when attending formal lessons and or discussions on the Singapore Cambridge GCE A Level Higher 2 H2 Chemistry As the Higher 1 H1 Chemistry syllabus is a subset of the H2 Chemistry syllabus this ebook is

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Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz,1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences

Study Guide for Mortimer's Chemistry, Sixth Edition Donald W. Shive,Louise E. Shive,1986

Chemistry for the IB Diploma Study and Revision Guide Christopher Talbot,Richard Harwood,2017-07-24 Exam Board IB Level IB Subject Chemistry First

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accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

Fundamentals of Chemistry, Study Guide James E. Brady,John R. Holum,1988-04-20 This Third Edition of

the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

General Chemistry Workbook Daniel C. Tofan,2010-07-28

This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students

a good idea of what they might expect on an actual final exam covering the entire material **Chemistry - Study Guide to Accompany Moore/Davies/Collins** Clare T. Furse,1978 *Study Guide to Accompany Chemical Principles, Properties, and Reactions* Kenneth L. Henold,1984 *TEAS 6 Test Prep Chemistry Review--Exambusters Flash Cards--Workbook 4 of 5*
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