

Phase Changes

Heating a substance in a given phase generally causes the temperature to increase. This causes the particles to move faster and collide harder. Harder, faster collisions cause particles to rebound harder moving them further apart. Larger distances between particles weakens the forces of attraction between them. When the forces of attraction are weak enough, the distance between the particles increases markedly and the phase changes. As a result, a solid melts and a liquid evaporates. The reverse happens when a substance cools. As the particles slow down and the collisions weaken, the particles move closer together increasing the forces of attraction between the particles. As a result, a gas condenses and a liquid freezes. At normal atmospheric pressures, some substances change directly from a solid to a gas. This change is called sublimation. This is what occurs when ice cubes shrink in the freezer, or when a wash dries on a clothesline on a cold winter day. Some substances, such as dry ice [$\text{CO}_2(\text{s})$], never pass through a liquid phase at standard pressure (101.3 kPa). A gas can also change directly to a solid. This is called deposition.



Answer the questions below based on your reading and on your knowledge of chemistry.

- In the space below are the labels solid, liquid, and gas. Draw arrows between them to represent phase changes. Label each arrow with the name of the phase change it represents. Two examples are done for you. Also, draw an arrow labeled temperature to show the direction in which temperature is increasing.



- Under what conditions does a solid melt? Why? When kinetic energy (temperature) rises above the forces of attraction, a solid melt
- Under what conditions does a gas condense? Why? When kinetic energy (temperature) decreases and the forces of attraction are felt by the particles

Evan Silberstein Chemistry Answer Key

J Elliott



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Constructing Worlds through Science Education John K. Gilbert, 2013-05-13 Internationally renowned and award winning author John Gilbert has spent the last thirty years researching thinking and writing about some of the central and enduring issues in science education He has contributed over twenty books and 400 articles to the field and is Editor in Chief of the International Journal of Science Education For the first time he brings together sixteen of his key writings in one volume This unique book highlights important shifts in emphasis in science education research the influence of important individuals and matters of national and international concern All this is interwoven in the following four themes explanation models and modeling in science education relating science education and technology education informal education in science and technology alternative conceptions and science education The Publishers' Trade List Annual ,1978 Whitaker's Cumulative Book List ,1977 **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain),1943 Includes list of members 1882 1902 proceedings of the annual meetings and various supplements British Chemical Abstracts ,1933 *Handbook of Non-prescription Drugs* ,2004 *British Chemical and Physiological Abstracts* ,1933 *British Abstracts* ,1947 **Reference Catalogue of Current Literature** ,1940 **Medical Books and Serials in Print** ,1984 *Chemistry Tests Answer Key (for Use With)* Bob Jones University Press, *Answer Key for Physical Setting Chemistry* ,2008-09-01 Answer Key for past New York State Regents in Physical Setting Chemistry *Chemistry* Antony C. Wilbraham,Dennis D. Staley,Michael S. Matta,Edward L. Waterman,2004-07-01 Prentice Hall Chemistry meets the needs of students with a range of abilities diversities and learning styles by providing real world connections to chemical concepts and processes The first nine chapters introduce students to the conceptual nature of chemistry before they encounter the more rigorous mathematical models and concepts in later chapters The technology backbone of the program is the widely praised Interactive Textbook with ChemASAP which provides frequent opportunities to practice and reinforce key concepts with Animations Simulations Assessment and Problem solving tutorials **Answer Key to Experimental Chemistry and Workbook** ,1969 Answer Key for General Chemistry Staff of Christ the King Books,2017-02 The Answer Key: A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success (Volume Two) (Preliminary E Rachel Turoscy,2017-12-31 **Chemistry Answer Key Units 6-10 (RES)** Responsive Educations Solutions Staff,2012-08-01 **The Answer Key: A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success (Volume One) (Preliminary E Rachel Turoscy,2017-12-31** Marthe Villalonga ,1978 Chemistry Answer Key Units 1-5 (RES) Responsive Education Solutions Staff,2011-08-01 Key Individual Answer Key for Chemistry Units

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