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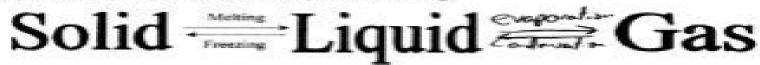
Phase Chariges

Heating a substance is a given phase generally causes the temperature to increase. This causes the particles to move flaster and coilide harder. Harder, firster collisions cause particles to rebound harder moving them flarther apart. Larger distances between particles weakers the forces of attraction however them. When the forces of attraction are weak enough, the distance between the particles increases markedly and the phase changes. As a result, a solid meths and a liquid evaporates. The reverse happens when a substance cools. As the particles slow down and the collisions weakers, the particles move closer together increasing the forces of attraction between the particles. As a result, a gas condenses and a liquid fractor. At normal immospheric pressures, some substances change directly from a solid to a gas. This change is coilled subtimation. This is what occurs when we cabes strink in the fractor, or when a wash dries on a clothesing on a cold winter day. Some substances, such as day ice [CO₂(x)], never pass through a liquid phase at mandard pressure [101.3 kPa]. A gas can also change directly to a solid. This is called deposition.



Answer the questions below based on your reading and on your knowledge of chemistry.

In the space below are the labels sollid, fiquid, and gas. Draw arrows between them to represent phase changes. Label each
arrow with the name of the phase change it represents. Two examples are done for you. Also, draw an arrow labeled
temperature to show the direction in which temperature is increasing.



- 2. Under what conditions does a mote many ways When Kinetic energy (temperature)

 Yister a have the forces of attraction, a solid milt
- decreases and the forces of attraction are foll by the fartiels

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