

Phase Changes

Heating a substance in a given phase generally causes the temperature to increase. This causes the particles to move faster and collide harder. Harder, faster collisions cause particles to rebound harder moving them further apart. Larger distances between particles weakens the forces of attraction between them. When the forces of attraction are weak enough, the distance between the particles increases markedly and the phase changes. As a result, a solid melts and a liquid evaporates. The reverse happens when a substance cools. As the particles slow down and the collisions weaken, the particles move closer together increasing the forces of attraction between the particles. As a result, a gas condenses and a liquid freezes. At normal atmospheric pressures, some substances change directly from a solid to a gas. This change is called sublimation. This is what occurs when ice cubes shrink in the freezer, or when a wash dries on a clothesline on a cold winter day. Some substances, such as dry ice [$\text{CO}_2(\text{s})$], never pass through a liquid phase at standard pressure (101.3 kPa). A gas can also change directly to a solid. This is called deposition.



Answer the questions below based on your reading and on your knowledge of chemistry.

- In the space below are the labels solid, liquid, and gas. Draw arrows between them to represent phase changes. Label each arrow with the name of the phase change it represents. Two examples are done for you. Also, draw an arrow labeled temperature to show the direction in which temperature is increasing.

Solid $\xrightleftharpoons[\text{Freezing}]{\text{Melting}}$ Liquid $\xrightleftharpoons[\text{Condensation}]{\text{Evaporation}}$ Gas

- Under what conditions does a solid melt? Why? When kinetic energy (temperature) rises above the forces of attraction, a solid melt
- Under what conditions does a gas condense? Why? When kinetic energy (temperature) decreases and the forces of attraction are felt by the particles

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