

Year 11 C2 - Discovering Chemistry Topic 1



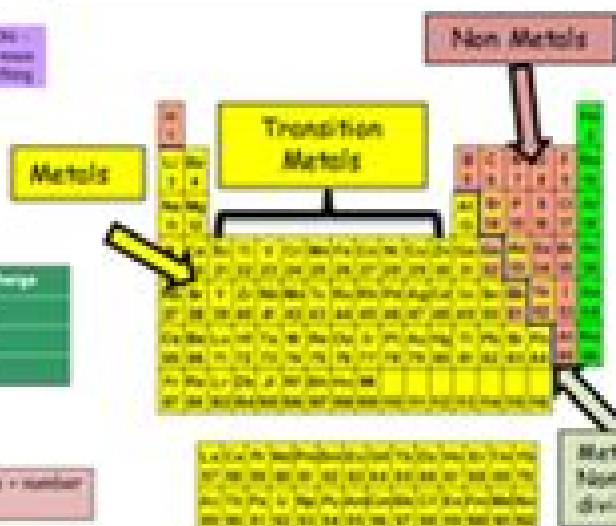
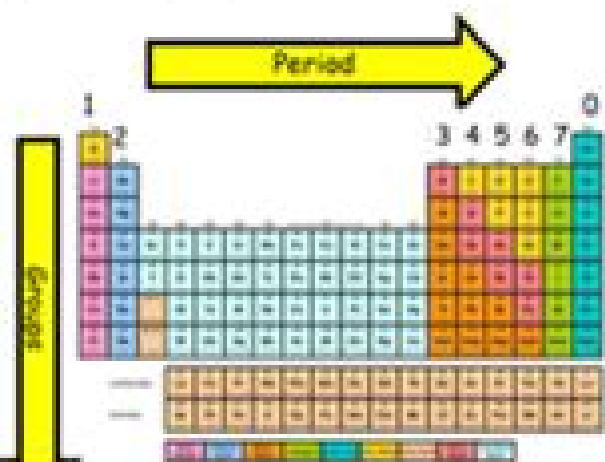
Particle	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-

He

4 ← **MASS NUMBER** = number of protons + number of neutrons

2 ← **PROTON NUMBER** = number of protons (atomic)

2 ← **CHARGE**



Isotopes

Arrange in a table with a different number of neutrons

Notice that the mass number is different. How many neutrons does each isotope have?

16	17	18
O	O	O
8	8	8

Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more.

How can an atom have a **decimal** for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

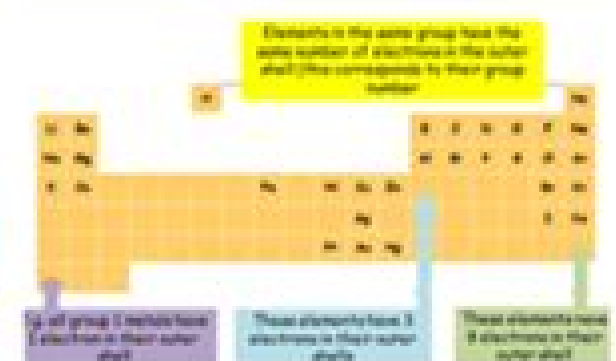
$35 \times 3 + 1 \times 37 = 35.5$

I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

I put elements with similar properties into horizontal rows. I also left gaps in case any new elements were discovered

Dmitri Mendeleev

<p>14</p> <p>N</p> <p>7</p> <p>Electron structure + 2, 5</p>	<p>24</p> <p>Mg</p> <p>12</p> <p>Electron structure + 2, 8, 2</p>	<p>40</p> <p>Ca</p> <p>20</p> <p>Electron structure + 2, 8, 8, 2</p>
---------------------------------------------------------------------------------------	--------------------------------------------------------------------------------------------	-----------------------------------------------------------------------------------------------



(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

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Richard Parsons

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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