Brønsted-Lowry Acid/Base Theory

- · Acids: Proton (H") donor
- · Bases: Proton (H') acceptor
- Conjugate Acids-Base Pair: Two substances that differ only by the presence of an H
 - Conjugate Acid: Formed when a proton is added to a base (positive charge)
 - Add a Hydrogen to the formula and add +1 charge
 - Conjugate Base: Formed when a proton is removed from an acid (negative charge)
 - Take away a Hydrogen from the formula and subtract -1 charge
- Amphoteric: The ability for a compound to act as an acid OR as a base (EX: H₂O)
- Acids are classified based on how many H* are present
 - Monoprotic: 1 H*
 - Diprotic: 2 H*
 - O Triprotic: 3 H'
 - Diprotic and triprotic acids are also amphoteric because they have more than one proton that they can lose

Properties of Strong and Weak Acids/Bases

Strong Acids/Bases	Weak Acids/Bases
Strong electrolytes Ionize completely in water Equilibrium is not present Very unlikely that the ions reform into products Very weak attraction between ions Split apart easily	Weak electrolytes Does not ionize completely in water Equilibrium exists Strong attraction between ions Do not spoilt apart easily
Strong acid	Weak acid
Weak attraction	Strong attraction Partial ionization

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David R. Klein, Laurie S. Starkey

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