

TYPES OF CHEMICAL REACTIONS LAB PART#1

I. Purpose: To view the actual chemical reactions, write the correct balanced chemical equation, and type of chemical reaction.

Connection to Business: Businesses utilize many types of chemical reactions to produce, maintain, and sterilize their product or to prevent pollution.

II. Materials: burner, test tube, metal tongs (only....no rubber ends), microspatula, evaporating dish, wood splints

Chemicals: copper (wire), magnesium (ribbon), copper (II) carbonate

SAFETY GOGGLES!

III. Procedure: Part #1

1. Obtain materials.
2. Use the fine sandpaper to clean a piece of copper wire until shiny.
3. Set up a Bunsen burner.
4. Using tongs, hold the wire in the hottest part of the burner flame for 1-2 minutes. Record your observation of the wire's appearance afterwards.
5. Place an evaporating dish near the Bunsen burner and repeat step #4 except place the remaining magnesium into the evaporating dish.
(WARNING: AVOID LOOKING DIRECTLY AT THE FLAME)
6. Place 2 heaping microspatulas of copper (II) carbonate into a clean, dry test tube.
7. Using a test tube holder, hold the test tube (at a 45° angle pointed away from people) over the Bunsen burner flame for 3 minutes or until completely black.
(Allow this to cool slightly before placing into the wood test tube holder.)
8. Light a wood splint and place inside the test tube. Record your observations.
9. Dispose of all the metals and burnt powders in a zip-lock bag and place in the trash.
10. Clean and return all lab equipment.

IV. Data/Questions: Part #1 Data Table

Sample	Appearance Before Reaction	Appearance After Reaction
1. Copper	<i>Answer: shiny, reddish brown</i>	<i>Answer: black, powder coating</i>
2. Magnesium	<i>Answer: silver, tinsel</i>	<i>Answer: black ashes, bright light</i>
3. Mercury (II) Oxide (will not perform experiment)	<i>Reddish powder</i>	<i>Liquid mercury formed; flame burned brighter when placed in the test tube</i>
4. Copper (II) Carbonate	<i>Answer: green powder</i>	<i>Answer: decomposes into gas CO₂ and black residue</i>

Describing Chemical Reactions Answer Key Guided Practice

Annelies Wilder-Smith



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