

Name _____

Section _____

Date _____

20. A gas sample occupies 200 mL at 700 mm Hg. What volume does the gas occupy at 600 mm Hg?

10.00

$$(200 \text{ mm Hg})(200 \text{ mL}) = (600 \text{ mm Hg}) V_2$$

$$V_2 = 333 \text{ mL}$$

21. Nitrogen gas in a sealed container exerts a pressure of 150 atm at 300 K. What will the pressure be if the temp. for the gas is 323 K, the volume and the amount of gas remain the same?

164

$$\frac{150 \text{ atm}}{300 \text{ K}} = \frac{P_2}{323 \text{ K}}$$

$$P_2 = 164 \text{ atm}$$

22. A sample of gas occupies a volume of 80 mL at a pressure of 1.50 atm and a temperature of 273 K. What will be the volume of a pressure of 1.50 atm and a temperature of 323 K?

31.5

$$\frac{(1.50 \text{ atm})(80 \text{ mL})}{273 \text{ K}} = \frac{(1.50 \text{ atm}) V_2}{323 \text{ K}}$$

$$V_2 = 31.55 \text{ mL}$$

Gas Laws Test Review Answers

RM Cervero



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