

Mixed Gas Laws Worksheet

- 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$\begin{aligned} \text{vol}^m &= 98 \text{ L} \\ P &= 2.8 \text{ atm} \\ t &= 292 \text{ K} \\ \text{mole} &= ? \\ R &= 0.082 \frac{\text{L atm}}{\text{mol K}} \end{aligned}$$

$$\begin{aligned} PV &= nRT \\ n &= \frac{PV}{RT} \\ &= \frac{2.8 \times 98}{0.082 \times 292} = \underline{\underline{11.46 \text{ mole}}} \end{aligned}$$

- 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25°C , what will the pressure of the resulting mixture of gases be?

$$\begin{aligned} &5 \text{ moles } \text{O}_2 \\ &3 \text{ moles } \text{N}_2 \\ \text{vol}^m &= 30 \text{ L} \\ \text{temp} &= 25^\circ\text{C} \rightarrow 273 + 25 \\ &= \underline{\underline{298 \text{ K}}} \\ P &= ? \end{aligned}$$

$$\begin{aligned} PV &= nRT \\ P &= \frac{nRT}{V} = \frac{8 \times 0.082 \times 298}{30 \text{ L}} \\ &= \underline{\underline{6.45 \text{ atm}}} \end{aligned}$$

- 3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0°C . By noon the temperature has risen to 45.0°C . What is the new volume of the balloon?

$$\begin{aligned} \textcircled{1} \text{ Vol}^m &= 35 \text{ L} \\ \textcircled{1} \text{ Temp} &= 20^\circ\text{C} = 293 \text{ K} \\ \textcircled{2} \text{ Temp} &= 45^\circ\text{C} = 318 \text{ K} \\ \textcircled{2} \text{ Vol}^m &= ? \end{aligned}$$

$$\begin{aligned} &\text{Charles law} \\ \frac{V_1}{T_1} &= \frac{V_2}{T_2} \\ \frac{35}{293} &= \frac{V_2}{318} \\ V_2 &= \frac{35 \times 318}{293} = \underline{\underline{37.78 \text{ L}}} \end{aligned}$$

- 4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$\begin{aligned} V &= 35 \text{ L} \\ T &= 315 \text{ K} \\ P &= 190 \text{ atm} \\ n &= ? \end{aligned}$$

$$\begin{aligned} PV &= nRT \\ n &= \frac{PV}{RT} = \frac{190 \times 35}{0.082 \times 315} \\ &= \underline{\underline{257.45 \text{ mol}}} \end{aligned}$$

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