



# COVALENT BONDING

## Practice Problems

Do your work on a separate sheet of graphing paper.

### SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each substance as ionic or molecular. Write its formula.

- a. NaCl (table salt)      b.  $\text{CH}_4$  (natural gas)      c.  $\text{H}_2\text{O}$  (water)  
 d.  $\text{KNO}_3$  (potassium nitrate)      e.  $\text{H}_2\text{SO}_4$  (sulfuric acid)

2. Write the formula for the following molecular compounds.

- a. carbon dioxide      b. carbon monoxide      c. dinitrogen  
 d. sulfur dioxide      e. sulfur trioxide      f. phosphorus

3. Write the formula for each compound. Write the name of each compound. Write the formula.

4. Write the formula for each compound. Write the name of each compound. Write the formula.

5. Write the formula for each compound. Write the name of each compound. Write the formula.

### SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride,  $\text{HF}$ .



2. Draw the Lewis structure for methane,  $\text{CH}_4$ .



3. Draw the Lewis structure for carbon dioxide,  $\text{CO}_2$ .



4. Draw the Lewis structure for acetylene,  $\text{C}_2\text{H}_2$ .

5. Draw the Lewis structure for each compound. Write the name of each compound. Write the formula.



6. Draw the Lewis structure for each compound. Write the name of each compound. Write the formula.

### SECTION 8.3 BONDING: THE ORBITAL

1. Write the Lewis structure for each compound. Write the name of each compound. Write the formula.

2. Write the Lewis structure for each compound. Write the name of each compound. Write the formula.

# Covalent Bonding Study Guide Answers Pearson Chemistry

**Padhraic Smyth**



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