

# Counting by Measuring Mass

## Purpose

To determine the mass of several samples of chemical compounds and use the data to count atoms

## Materials

H<sub>2</sub>O, NaCl, and CaCO<sub>3</sub>  
plastic spoon  
weighing paper  
balance

## Safety



## Procedure

Measure the mass of one level teaspoon of water (H<sub>2</sub>O), sodium chloride (NaCl), and calcium carbonate (CaCO<sub>3</sub>). Make a table similar to the one below.

	H <sub>2</sub> O(l)	CaCl <sub>2</sub>	Na <sub>2</sub> CO <sub>3</sub>
Mass (g)	2.42 g	3.95 g	2.30 g
Molar mass (g/mol)	18.0 g/mol	111.1 g/mol	106.0 g/mol
Moles of each compound	0.146 mol	0.0354 mol	0.0217 mol
Moles of each element	0.292 mol H 0.146 mol O	0.0354 mol Ca 0.0708 mol Cl	0.0434 mol Na 0.0217 mol C 0.0651 mol O
Atoms of each element	1.76 × 10 <sup>23</sup> H 1.79 × 10 <sup>23</sup> O	2.15 × 10 <sup>23</sup> Ca 4.26 × 10 <sup>23</sup> Cl	2.61 × 10 <sup>23</sup> Na 1.31 × 10 <sup>23</sup> C 3.92 × 10 <sup>23</sup> O



$$2 \times 23.0 + 12.0 + 3 \times 16.0 = 106.0$$

$$2.30 \text{ g} \left( \frac{1 \text{ mol}}{106.0 \text{ g}} \right) = 0.0217$$

$$2.30 \text{ g} \left( \frac{1 \text{ mol}}{106.0 \text{ g}} \right)$$

1. **Calculate** Determine the number of moles of H<sub>2</sub>O contained in one level teaspoon.

$$\text{moles of H}_2\text{O} = \text{g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}}$$

Repeat for the remaining compounds. Use the periodic table to calculate the molar masses of NaCl and CaCO<sub>3</sub>.

2. **Calculate** Determine the number of moles of each element present in the teaspoon-sized sample of H<sub>2</sub>O.

$$\text{moles of H} = \text{mol H}_2\text{O} \times \frac{2 \text{ mol H}}{1 \text{ mol H}_2\text{O}}$$

Repeat for the other compounds in your table.

3. **Calculate** Determine the number of atoms of each element present in the teaspoon-sized sample of H<sub>2</sub>O.

$$\text{atoms of H} = \text{mol H} \times \frac{6.02 \times 10^{23} \text{ atoms H}}{1 \text{ mol H}}$$

Repeat for the other compounds in your table.

# Counting By Measuring Mass Lab Answer Key

**L Reisser**



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