



Dry Lab 3

Atomic and Molecular Structure

Metallic cations heated to high temperatures produce characteristic colors that appear in the starbursts.

- To view and calibrate visible line spectra
- To identify an element from its visible line spectrum
- To identify a compound from its infrared spectrum
- To predict the three-dimensional structure of molecules and molecular ions

OBJECTIVES

Visible light, as we know it, is responsible for the colors of nature—blue skies, green trees, red roses, orange-red rocks, and brown deer. Our eyes are able to sense and distinguish the subtleties and the intensities of those colors through a complex, naturally designed detection system—our eyes. The beauty of nature is a result of the interaction of sunlight with matter. Since all matter consists of atoms and molecules, it is obvious that sunlight, in some way, interacts with them to produce nature's colors.

Within molecules of compounds are electrons, vibrating bonds, and rotating atoms, all of which can absorb energy. Because every compound is different, every molecule of a given compound possesses its own unique set of electronic, vibrational, and rotational **energy states**, which are said to be **quantized**. When incident electromagnetic (EM) radiation falls on a *molecule*, the radiation absorbed (the absorbed light) is an energy equal to the difference between two energy states, placing the molecule in an excited state. The remainder of the EM radiation passes through the molecule unaffected.

Atoms of elements interact with EM radiation in much the same way, except there are no bonds to vibrate or atoms to rotate. Only electronic energy states are available for energy absorption.

EM radiation is energy as well as light, and light has wavelengths and frequencies. The relationship between the energy, E , and its wavelength, λ , and frequency, ν , is expressed by the equations:

$$E = \frac{hc}{\lambda} = h\nu \quad (\text{D3.1})$$

where h is Planck's constant, 6.63×10^{-34} J·s/photon; c equals the velocity of the EM radiation, 3.00×10^8 m/s; λ is the wavelength of the EM radiation in meters; and ν (pronounced "new") is its frequency in reciprocal seconds (s^{-1}).

EM radiation includes not only the wavelengths of the visible region (400 to 700 nm) but also those that are shorter (e.g., the ultraviolet and X-ray regions) and longer (e.g., the infrared, microwave, and radio wave regions). See Figure D3.1.

From Equation D3.1, shorter wavelength EM radiation has higher energy. When an atom or molecule absorbs EM radiation from the visible region of the spectrum that is excited from a lower to a higher energy state.

INTRODUCTION

Energy state: the amount of energy confined within an atom or molecule, which can be changed by the absorption or emission of discrete (quantized) amounts of energy

Quantized: only a definitive amount (of energy)

Wavelength: the distance between two crests of a wave

Frequency: the number of crests that pass a given point per second

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Table of Contents Dry Lab 3 Atomic And Molecular Structure

1. Understanding the eBook Dry Lab 3 Atomic And Molecular Structure
 - The Rise of Digital Reading Dry Lab 3 Atomic And Molecular Structure
 - Advantages of eBooks Over Traditional Books
2. Identifying Dry Lab 3 Atomic And Molecular Structure
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Dry Lab 3 Atomic And Molecular Structure
 - User-Friendly Interface
4. Exploring eBook Recommendations from Dry Lab 3 Atomic And Molecular Structure

- Personalized Recommendations
- Dry Lab 3 Atomic And Molecular Structure User Reviews and Ratings
- Dry Lab 3 Atomic And Molecular Structure and Bestseller Lists
- 5. Accessing Dry Lab 3 Atomic And Molecular Structure Free and Paid eBooks
 - Dry Lab 3 Atomic And Molecular Structure Public Domain eBooks
 - Dry Lab 3 Atomic And Molecular Structure eBook Subscription Services
 - Dry Lab 3 Atomic And Molecular Structure Budget-Friendly Options
- 6. Navigating Dry Lab 3 Atomic And Molecular Structure eBook Formats
 - ePub, PDF, MOBI, and More
 - Dry Lab 3 Atomic And Molecular Structure Compatibility with Devices
 - Dry Lab 3 Atomic And Molecular Structure Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Dry Lab 3 Atomic And Molecular Structure
 - Highlighting and Note-Taking Dry Lab 3 Atomic And Molecular Structure
 - Interactive Elements Dry Lab 3 Atomic And Molecular Structure
- 8. Staying Engaged with Dry Lab 3 Atomic And Molecular Structure
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Dry Lab 3 Atomic And Molecular Structure
- 9. Balancing eBooks and Physical Books Dry Lab 3 Atomic And Molecular Structure
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Dry Lab 3 Atomic And Molecular Structure
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Dry Lab 3 Atomic And Molecular Structure
 - Setting Reading Goals Dry Lab 3 Atomic And Molecular Structure
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Dry Lab 3 Atomic And Molecular Structure

- Fact-Checking eBook Content of Dry Lab 3 Atomic And Molecular Structure
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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