

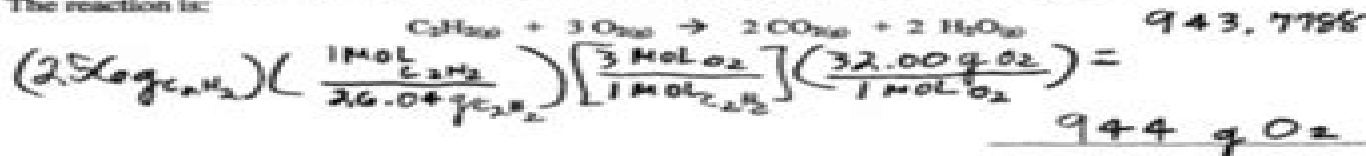
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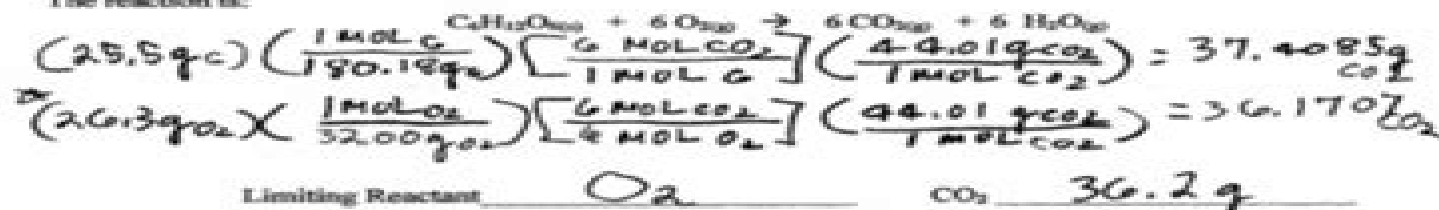
Please show problem set-ups, use correct units, and round to the correct number of significant figures for all calculations to receive full credit. Use scientific notation for answers when appropriate. PRINT CLEARLY all answers on the lines provided. Constants: $R = 0.0821 \text{ L atm/mol K}$

Molar Mass (g/mol): C= 12.01 O = 16.00 H= 1.01 N= 14.01 Na= 22.99 K= 39.10 Al= 26.98 Cl= 35.45

1. How many grams of oxygen gas are required to burn 256 g of acetylene gas (C_2H_2)?
The reaction is:



2. How many grams of carbon dioxide can be formed if 25.5 grams of glucose ($\text{C}_6\text{H}_{12}\text{O}_6$) is oxidized with 26.3 grams of oxygen gas? Also, show the calculation to determine the limiting reactant.
The reaction is:



3. Calculate the theoretical yield for a reaction if the actual yield is 555 grams and the reaction has a percent yield of 85.5 %.

$$\% \text{YIELD} = \left(\frac{\text{ACTUAL}}{\text{THEO}} \right) (100\%) \quad \text{THEO} = 649.12 \text{ g}$$

$$(85.5\%) = \frac{555 \text{ g} (100\%)}{\text{THEO}} \quad \underline{649 \text{ g}} \quad -2 \text{ DEC}$$

4. Calculate the molarity of a solution of potassium chloride that has a volume of 255 mL and contains 13.78 grams of potassium chloride.

$$M = \frac{\text{g/mol}}{\text{VOL}} = \frac{13.78 \text{ g}}{74.55 \text{ g/mol}} = 0.18497 \text{ mol/L}$$

$$\underline{0.185 \text{ M}}$$

5. How many grams of sodium hydroxide are required to make 2.555 L of 0.155 M solution?

$$M = \frac{\text{g/mol}}{\text{VOL}} \quad 0.155 \frac{\text{mol}}{\text{L}} = \frac{\text{g/40.00 g/mol}}{2.555 \text{ L}} \quad \underline{15.8 \text{ g}}$$

$$\text{g} = 15.841$$

Catalyst Lab Manual Chem 101 Answer Solutions

Thomas Griffiths



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