

On each of the problems below:

- Draw particle diagrams of the initial and final conditions.
- Organize **given** information (P, V, T, n) with Initial/Final/Effect Table.
- Make a decision as to how a change in P, V, T, or n will affect the starting quantity (show with arrows up or down in Effect line – example:  $P \uparrow V \downarrow$  - which means if the pressure increases, the volume decreases)
- Setup conversion equation by multiplying by the appropriate factors.

1. You collect a sample of gas that occupies 150 mL at 25 °C. What is its volume when you increase the temperature to 50°C? (P and n = constant)

	P	T	V	n
Initial				
Final				
Effect				

2. What would be the new pressure if 250 cm<sup>3</sup> of gas at standard pressure is compressed to a volume of 150 cm<sup>3</sup>? (\_\_\_\_\_ = constant)
3. What would be the new volume if 250 cm<sup>3</sup> of gas at 25°C and 730 mm pressure were changed to standard conditions of temperature and pressure? (\_\_\_\_\_ = constant)
4. Your bike tire contains 15 units of air particles, has a volume of 160mL, and has a pressure of 13 psi. Your tire develops a leak. Now it contains 10 units of air and has contracted to a volume of 150mL. What is the tire pressure now?

# Chemistry Unit 2 Pvtn Problems Answers

**Didier Musso**



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