

7. GROUPS 1 & 2

SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 1 Magnesium oxide, known as magnesia, can be produced by the thermal decomposition of magnesium carbonate.

The equation for the reaction is



- *(a) Other Group 2 carbonates also undergo thermal decomposition.

State and explain the trend in thermal stability of carbonates down Group 2.

(4)

Thermal stability increases down the group

Because ionic radius increases as you go down the group

This results in less polarisation (distortion) of the carbonate ion

And less weakening of the covalent bonds (C-O) within the carbonate ion

- *(c) State why carrying out a flame test is **not** an appropriate way to confirm the presence of magnesium ions in the sample of magnesia.

(1)

Magnesium ions give no flame colour

OR

Energy is emitted with Mg ions is outside of the visible region

(Total for Question 1 = 15 marks)

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Léo Pomar



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Table of Contents Edexcel Biology B3 Past Papers Mark Scheme

1. Understanding the eBook Edexcel Biology B3 Past Papers Mark Scheme
 - The Rise of Digital Reading Edexcel Biology B3 Past Papers Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying Edexcel Biology B3 Past Papers Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Edexcel Biology B3 Past Papers Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from Edexcel Biology B3 Past Papers Mark Scheme
 - Personalized Recommendations
 - Edexcel Biology B3 Past Papers Mark Scheme User Reviews and Ratings
 - Edexcel Biology B3 Past Papers Mark Scheme and Bestseller Lists
5. Accessing Edexcel Biology B3 Past Papers Mark Scheme Free and Paid eBooks

- Edexcel Biology B3 Past Papers Mark Scheme Public Domain eBooks
 - Edexcel Biology B3 Past Papers Mark Scheme eBook Subscription Services
 - Edexcel Biology B3 Past Papers Mark Scheme Budget-Friendly Options
6. Navigating Edexcel Biology B3 Past Papers Mark Scheme eBook Formats
- ePub, PDF, MOBI, and More
 - Edexcel Biology B3 Past Papers Mark Scheme Compatibility with Devices
 - Edexcel Biology B3 Past Papers Mark Scheme Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Edexcel Biology B3 Past Papers Mark Scheme
 - Highlighting and Note-Taking Edexcel Biology B3 Past Papers Mark Scheme
 - Interactive Elements Edexcel Biology B3 Past Papers Mark Scheme
8. Staying Engaged with Edexcel Biology B3 Past Papers Mark Scheme
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Edexcel Biology B3 Past Papers Mark Scheme
9. Balancing eBooks and Physical Books Edexcel Biology B3 Past Papers Mark Scheme
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Edexcel Biology B3 Past Papers Mark Scheme
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Edexcel Biology B3 Past Papers Mark Scheme
- Setting Reading Goals Edexcel Biology B3 Past Papers Mark Scheme
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Edexcel Biology B3 Past Papers Mark Scheme
- Fact-Checking eBook Content of Edexcel Biology B3 Past Papers Mark Scheme
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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