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Your success in chemistry depends on your ability to solve numerical problems.

BASIC SKILLS

CALCULATOR SURVIVAL

- Become familiar with your calculator.
- Before the exam, make sure you can multiply, divide, add, subtract and use all needed functions.
- Calculators never make mistakes; they take your input (intended or accidental) and give an answer.
- Look at the answer; does it make sense?
- Do a quick estimate to check your work.
- **Sample:** $4.34 \times 7.88 \div 1.55 \times 9.81$ is roughly $4 \times 8 \div 1.5 \times 10$ or 3.2; the actual answer, 3.24

HOW TO DO WORD PROBLEMS

- Read and evaluate the question before you start plugging numbers into the calculator.
- Identify the variables, constants and equations.
- Write out units of the variables and constants.
- Work out the unit before the number.
- You may have extra information, or you may need to obtain constants from your text.

HOW TO WORK WITH UNITS

All numerical data has units. In chemistry we use metric "SI" units.

Pitfall: If the unit is wrong, the answer is wrong!

1. Unit prefixes: denote powers of 10^3

tera T 10^{12}	giga G 10^9	mega M 10^6
kilo k 10^3	deci d 10^{-1}	centi c 10^{-2}
milli m 10^{-3}	micro μ 10^{-6}	nano n 10^{-9}
pico p 10^{-12}	femto f 10^{-15}	

2. Fundamental Units

- Mass: kilogram (kg)
- Length: meter (m)
- Temperature: kelvin (K)
- Time: second (s)
- Amount of a substance: mole
- Electric charge: coulomb (C)

3. Derived Units

- Area: length squared, m^2 ; 1 liter (L) = 1 dm^3
- Volume: length cubed, m^3 ; 1 liter (L) = 1 dm^3
- Density: mass/volume; common unit kg/m^3
- Speed: distance/time; common unit m/s
- Electric current: ampere (A) = 1 C/s
- Force: Newton (N) = 1 kg m/s^2
- Energy: joule (J) = 1 kg m^2/s^2
- Pressure: Pascal (Pa) = 1 $kg/m \cdot s^2$

4. Fundamental Constants

	In units (symbol)	In kg
Avogadro's number, N_A	6.022×10^{23}	6.022×10^{23}
Ideal gas constant, R	8.314 J $K^{-1} mol^{-1}$	8.314 J $K^{-1} mol^{-1}$
R, for gas calculations	0.0821 L $atm / K \cdot mol$	0.0821 L $atm / K \cdot mol$
R, for energy calculations	8.314 J $K^{-1} mol^{-1}$	8.314 J $K^{-1} mol^{-1}$
Faraday's Constant, F	96,485 C/mol	96,485 C/mol
Planck's Constant, h	$6.626 \times 10^{-34} J \cdot s$	$6.626 \times 10^{-34} J \cdot s$
Speed of Light, c	$2.9979 \times 10^8 m/s$	$2.9979 \times 10^8 m/s$

BASIC SKILLS cont.

HOW TO CONVERT DATA

The unit & numerical value are changed using a conversion factor or equation.

HOW TO USE EQUATIONS FOR DATA CONVERSION

- **Sample:** Convert a temperature of 43°C to °F and Kelvin.
Given: $K = ^\circ C + 273.15$
Calculate: Temperature in K = $43^\circ C + 273.15 = 316.15 K$
Given: $^\circ F = 9/5 \cdot ^\circ C + 32$
Calculate: Temperature in $^\circ F = 9/5 \times 43^\circ C + 32 = 111.4^\circ F$



HOW TO USE A CONVERSION FACTOR

- **Sample:** "1 hour = 60 minutes" gives two conversion factors:
1. Multiply by "1 hour/60 minutes" to convert minutes to hours.
2. Multiply by "60 minutes/1 hour" to convert hours to minutes.
- **Sample:** Determine the number of hours in 43 minutes.
Given: Conversion factor is "1 hour/60 minutes."
Calculate: Time = 43 minutes \times (1 hour/60 minutes) = 0.717 hour (calculator cancels)

Common conversions:

1 caloric = 4.184 J	1 kg = 2.2 lb	1 m = 1.1 yd
1 gal = 3.785 L	1 angstrom (Å) = $1 \times 10^{-10} m$	
1 atm = 1.01325 $\times 10^5$ Pa	1 mm = 10 mm Hg	

Pitfall: Equations and conversion factors have units. Be sure of inverting conversion factors.

MATH REVIEW

HOW TO WORK WITH ALGEBRAIC EQUATIONS

Give equal treatment to each side.

- Add or subtract:
• **Sample:** Given, $x + y = 5$, then, $4 + x = 4 + 5$
- Multiply or divide:
• **Sample:** Given, $x = y$, then, $4x = 4y$ and $x/2 = y/2$
Given, $ax = b$, then $ax/2 = b/2$ and $bx/2a = x/2$

Pitfall: Dividing by zero is not allowed.

HOW TO WORK WITH SCIENTIFIC NOTATION

The exponent gives the power of 10.

- **Sample:** $0.00045 = 4.5 \times 10^{-4}$; $0.45 = 4.5 \times 10^{-1}$
- **Sample chemical applications:** Molecular diameters are 10^{-10} m; 1 liter of water contains about 1×10^{25} atoms.

MATH REVIEW cont.

HOW TO USE LOGARITHMS & EXPONENTS

- Common logarithm, \log_{10} base "10".
Denotes number or function in powers of 10.
• **Sample:** Given, $y = 10^5$, then $\log_{10} y = 5$.
- Natural logarithm, \ln , base " e ".
Denotes number or function in powers of " e ".
• **Sample:** Given, $x = e^2$, then $\ln x = 2$.
- Products: $\log(xy) = \log x + \log y$
- Powers: $\log(x^y) = y \log x$
- Multiplication: Add exponents from each term.
• **Sample:** $10^3 \times 10^2 = 10^{3+2} = 10^5$
- Division: Subtract denominator exponents from numerator exponents.
• **Sample:** $10^5/10^2 = 10^{5-2} = 10^3$
- Square root: $\sqrt{x} = x^{1/2}$
- Reciprocal: $1/x = x^{-1}$

Pitfall: Your calculator has separate keys for \ln , \log , e^x , 10^x and x^y .
• **Sample chemical applications:** pH of acid and base.

HOW TO CALCULATE ROOTS OF A POLYNOMIAL

An equation of the form $ax^2 + bx + c = 0$, has 2 solutions or roots, given by the quadratic formula:

$$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$$

- **Sample:** Determine the roots for the equation: $3x^2 + 4x + 1 = 0$

$$\text{Given: } a = 3, b = 4, c = 1$$

$$\text{Calculate: Root 1} = \frac{-4 + \sqrt{4^2 - 4 \times 3 \times 1}}{2 \times 3} = \frac{-4 + \sqrt{16 - 12}}{6} = \frac{-4 + 2}{6} = -\frac{2}{6} = -\frac{1}{3}$$

$$\text{Calculate: Root 2} = \frac{-4 - \sqrt{4^2 - 4 \times 3 \times 1}}{2 \times 3} = \frac{-4 - 2}{6} = -\frac{6}{6} = -1$$

Pitfall: Always round-off only. Substitute the roots into the equation to verify results.
• **Sample chemical applications:** Weak acids, weak bases, buffers, chemical equilibrium

HOW TO DETERMINE THE EQUATION OF A LINE

LINEAR EQUATION: $y = mx + b$

- m : slope of the line; $m = \Delta y/\Delta x$
- b : y intercept, the line crosses the y-axis at $(0, b)$ or $y = -mx$.
- **Sample:** Determine the equation of a line using (x,y) data.
Given: $(-2, -4)$, $(0, -2)$, $(2, 0)$
 $y = -2x - 4$, $y = -2x - 2$, $y = -2x - 0$
- Calculate: Slope = $m = \Delta y/\Delta x$
 $= \frac{12 - (-20)}{2 - (-2)} = \frac{32}{4} = 8$
- Calculate: y-intercept = $b = y_1 - mx_1$
 $= -4 - 2 \times (-2) = -4 + 4 = 0$

The equation of the line is $y = 8x + 4$.
• **Sample chemical applications:** Gas calculations; Beer's Law; analyzing reaction-rate data



Chemistry Study Guide Questions And Concepts

Terry C. Jones



Chemistry Study Guide Questions And Concepts:

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

Robinson Chemistry Study Guide Robinson, 1992

Organic Chemistry Michael B. Smith, 2016-03-09 Based on the premise that many if not most reactions in organic chemistry can be explained by variations of fundamental acid base concepts Organic Chemistry An Acid Base Approach provides a framework for understanding the subject that goes beyond mere memorization Using several techniques to develop a relational understanding it helps students fully grasp the essential concepts at the root of organic chemistry This new edition was rewritten largely with the feedback of students in mind and is also based on the author's classroom experiences using the first edition Highlights of the Second Edition Include Reorganized chapters that improve the presentation of material Coverage of new topics such as green chemistry Adding photographs to the lectures to illustrate and emphasize important concepts A downloadable solutions manual The second edition of Organic Chemistry An Acid Base Approach constitutes a significant improvement upon a unique introductory technique to organic chemistry The reactions and mechanisms it covers are the most fundamental concepts in organic chemistry that are applied to industry biological

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Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs **Chemistry** John Kenkel, Paul B. Kelter, David S. Hage, 2000-09-21

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