

## Exam I Study Guide

- **Physical Changes vs. Chemical Changes (Can you undo it?)**
  - Physical changes result in changes of appearance, but not composition.
    - Melting, evaporation, freezing, etc.
  - Chemical changes result in changes of composition of matter. (Chemical reaction)
    - Burning, mixing, dissolving, etc.
- **Physical Properties vs. Chemical Properties**
  - Physical properties can be described by observing the matter without a change in the composition.
    - Color, odor, density, temperature, etc.
  - Chemical properties can only be detected by altering the composition.
    - Flammability, oxidization, etc.
- **Atoms vs. Molecules**
  - Atoms are the smallest unit of a given element that carry the properties of that element.
  - Molecules are composed of two or more atoms.
- **Elements vs. Compounds**
  - Elements occur in nature and cannot be broken down to simpler particles through ordinary means.
  - Compounds are the combination of two or more elements.
- **Accuracy vs. Precision**
  - Accuracy refers to the agreement of a measurement with a known or accepted value. (Bullseye)
    - Correctness
  - Precision refers to the agreement of a measurement with other measurements in a given set of measurements. (Missing the bullseye but landing all darts in one place)
    - Repeatability or reproducibility
- **Rules Concerning Significant Figures**
  - All nonzero digits ARE significant.
  - Zeros between nonzero digits ARE significant.
  - Zeros at the beginning of a number are NOT significant; they only indicate the position of the decimal point.
  - Zeros at the end of a number to the right of a decimal point ARE significant.
  - Zeros at the end of a number with no decimal point are ASSUMED to be NOT significant, UNLESS we indicate their significance by placing a line over the last zero that is significant.

# General Chemistry Final Exam Study Guide

**SB Merriam**



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## **Table of Contents General Chemistry Final Exam Study Guide**

1. Understanding the eBook General Chemistry Final Exam Study Guide
  - The Rise of Digital Reading General Chemistry Final Exam Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Final Exam Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an General Chemistry Final Exam Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Final Exam Study Guide
  - Personalized Recommendations
  - General Chemistry Final Exam Study Guide User Reviews and Ratings
  - General Chemistry Final Exam Study Guide and Bestseller Lists
5. Accessing General Chemistry Final Exam Study Guide Free and Paid eBooks
  - General Chemistry Final Exam Study Guide Public Domain eBooks
  - General Chemistry Final Exam Study Guide eBook Subscription Services
  - General Chemistry Final Exam Study Guide Budget-Friendly Options

6. Navigating General Chemistry Final Exam Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - General Chemistry Final Exam Study Guide Compatibility with Devices
  - General Chemistry Final Exam Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of General Chemistry Final Exam Study Guide
  - Highlighting and Note-Taking General Chemistry Final Exam Study Guide
  - Interactive Elements General Chemistry Final Exam Study Guide
8. Staying Engaged with General Chemistry Final Exam Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers General Chemistry Final Exam Study Guide
9. Balancing eBooks and Physical Books General Chemistry Final Exam Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection General Chemistry Final Exam Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Final Exam Study Guide
  - Setting Reading Goals General Chemistry Final Exam Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Final Exam Study Guide
  - Fact-Checking eBook Content of General Chemistry Final Exam Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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- Interactive and Gamified eBooks

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