

IDEAL (2)

$$(8) \quad V = \frac{nRT}{P} = \frac{(225 \text{ mol})(.0821)(298 \text{ K})}{(1.05 \text{ atm})} = \boxed{5240 \text{ L}}$$

calc: 5242.67 L

$$(9) \quad n = \frac{PV}{RT} = \frac{(1.03 \text{ atm})(32.5 \text{ L})}{(.0821)(305 \text{ K})} = \boxed{1.34 \text{ mol}} \times \frac{44.0 \text{ g CO}_2}{1 \text{ mol}}$$

$$= \boxed{58.8 \text{ g CO}_2}$$

$$(10) \quad n = \frac{PV}{RT} = \frac{(.997 \text{ atm})(58.4 \text{ L})}{(.0821)(275.5 \text{ K})} = \boxed{2.57 \text{ mol}}$$

$$\text{molar mass} = \frac{10.3 \text{ g}}{2.57 \text{ mol}} = \boxed{4.00 \text{ g/mol}}$$

Helium

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Yan Bai



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