

The rate of a reaction depends on the concentration of the reactants. As the reaction proceeds to forward direction, concentration of reactant decreases and that of product increases. So, the rate of reaction generally decreases during the course of reaction.

3. Prove that decomposition of  $\text{H}_2\text{O}_2$  is a bimolecular reaction.



It is experimentally found that the reaction is **first order** with respect to both  $\text{H}_2\text{O}_2$  and  $\text{I}^-$ , which indicates that  $\text{I}^-$  is also involved in the reaction.

**Mechanism :**

**Step I :**



concentration of the reactants (or) increase in the concentration of products.	of the reactants is unity.
--	----------------------------

3. It depends on the initial concentration of reactants.	It does not depend on the initial concentration of reactants.
--	---

5. Consider the oxidation of nitric oxide to form  $\text{NO}_2$ .



a) Express the rate of the reaction in terms of changes in the concentration of  $\text{NO}$ ,  $\text{O}_2$  and  $\text{NO}_2$ .

**Soln :**

$$\text{Rate} = -\frac{1}{2} \frac{d[\text{NO}]}{dt} = -\frac{d[\text{O}_2]}{dt} = \frac{1}{2} \frac{d[\text{NO}_2]}{dt}$$

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