

## Study Guide Unit 9

1. KNOW the following vocabulary words (this means you NEED to be able to DEFINE them ALL):

a. Solution-

m. Solute-

b. Homogenous-

n. Solvent-

c. Heterogeneous-

o. Acid-

d. Soluble-

p. Base-

e. Insoluble-

q. Titration-

f. Miscible-

r. Analyte-

g. Immiscible-

s. Titrant-

h. Solubility-

t. Monoprotic acid-

i. Alloy-

u. Diprotic Acid

j. Saturated-

v. Triprotic Acid

k. Unsaturated-

l. Supersaturated-

2. How is the solubility of a gas affected by temp. and pressure?

Pressure up = solubility up

Temp up = solubility down

3. Under what conditions are salts most soluble?

High temperatures

4. Use the solubility curve to the right to answer the following.

a. What mass of  $\text{KNO}_3$  will make a saturated solution at  $60^\circ\text{C}$ ? 100g

b. Which of the compounds shown on the solubility curve decrease solubility as temperature increases?

$\text{NH}_3$ ,  $\text{Ce}_2(\text{SO}_4)_3$

c. What type of solution do I have if I **add** 110 g of  $\text{NaNO}_3$  in 100g of  $20^\circ\text{C}$  water?

Saturated

d. What mass of  $\text{KCl}$  can go into solution at  $80^\circ\text{C}$ ?

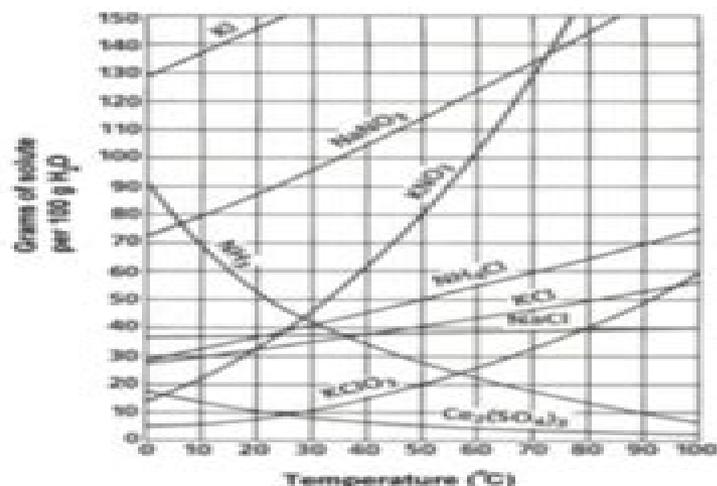
50 g

e. What is the least soluble compound at  $20^\circ\text{C}$ ?

$\text{KClO}_3$

f. Excluding  $\text{KI}$ , what is the most soluble compound at  $80^\circ\text{C}$ ?

$\text{KNO}_3$



# Chemistry Study Guide Answers What Are Solutions

**Leo J. Malone, Theodore O. Dolter**



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