

Balance the equation below



27) Once the equation is balanced, what is the **molar ratio** of water (H<sub>2</sub>O) to oxygen (O<sub>2</sub>)? 1:1

28) If 4.59 grams of O<sub>2</sub> react, how many grams of CH<sub>4</sub> are needed?

$$\frac{4.59 \text{ g-O}_2}{1} \times \frac{1 \text{ mole-O}_2}{32.00 \text{ g-O}_2} \times \frac{1 \text{ mole-CH}_4}{2 \text{ moles-O}_2} \times \frac{16.05 \text{ g CH}_4}{1 \text{ mole-CH}_4}$$

Answer: 1.15 g CH<sub>4</sub>

29) How many grams of CO<sub>2</sub> are needed to react with 8.95 grams of H<sub>2</sub>O?

$$\frac{8.95 \text{ g-H}_2\text{O}}{1} \times \frac{1 \text{ mole-H}_2\text{O}}{18.02 \text{ g-H}_2\text{O}} \times \frac{1 \text{ mole-CO}_2}{2 \text{ moles-H}_2\text{O}} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mole-CO}_2}$$

Answer: 10.93 g CO<sub>2</sub>

Balance the equation below



30) Once the equation is balanced, what is the **molar ratio** of H<sub>2</sub> to O<sub>2</sub>? 2:1

31) How many moles of H<sub>2</sub> are needed to react with 0.42 moles of H<sub>2</sub>O?

$$\frac{0.42 \text{ moles-H}_2\text{O}}{1} \times \frac{2 \text{ moles H}_2}{2 \text{ moles-H}_2\text{O}}$$

Answer: 0.42 moles H<sub>2</sub>

32) How many grams of H<sub>2</sub>O are needed to react with 4.55 moles of H<sub>2</sub>?

$$\frac{4.55 \text{ moles-H}_2}{1} \times \frac{2 \text{ moles-H}_2\text{O}}{2 \text{ moles-H}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mole-H}_2\text{O}}$$

Answer: 81.99 g H<sub>2</sub>O

33) How many moles of H<sub>2</sub>O are produced when reacted with 8.35 grams of H<sub>2</sub>?

$$\frac{8.35 \text{ g-H}_2}{1} \times \frac{1 \text{ mole-H}_2}{2.02 \text{ g-H}_2} \times \frac{2 \text{ moles H}_2\text{O}}{2 \text{ moles-H}_2}$$

Answer: 4.13 moles H<sub>2</sub>O

# Chemistry Study Guide Answers For Stoichiometry

**Cybellium**



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*The Practice of Chemistry Study Guide & Solutions Manual* Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences

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