Chapter 19 Study Guide: Acids, Bases & Titration - Answers

Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	ACIDS	BASES	
Arrhenius	Form hydronium ions (H ₂ O*) in aqueous sol'n	Form hydroxide ions (OH·) in aqueous sol'n	
Brønsted-Lowry	Proton (H*) donors	Proton (H*) acceptors	
Lewis	Electron pair acceptors	Electron pair donors	

Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:

	Α	B	CB.	CA
21	H2CO3 +	CaCl ₂ →	CarCO ₃ +	2HCI
	A	В	CB	CA
Ib.	2 HCl +	Ca(OH)z	CaCla	+ 2 H ₂ O

Write the names or formulas for the following acids and bases:

a.	H ₂ S - hydrosulfuric acid
b.	HC2H3O2-acetic acid
C.	NaOH - sodium hydroxide
di	HNO: - nitric acid

- e. Nitrous acid HNO₂
- f. Calcium hydroxide Ca(OH)₂
- g hydrochloric acid HCl
- Use the following formulas for the remaining questions: (these will <u>not</u> be given to you on the test!)

$$K_w = [H_3O^*][OH^*] = 1.0 \times 10^{-14}$$

 $pH = -log[H_3O^*]$
 $pOH = -log[OH^*]$

$$pH + pOH = 14$$

$$M_A \cdot V_A \cdot n_A = M_B \cdot V_B \cdot n_B$$

4. What is the pH of 1.34 x 10⁻⁴ M solution of hydrochloric acid?

$$pH = -log [1.34 \times 10^{-4}M] = 3.873$$

Determine the pH and pOH of a 1.85 x 10⁻⁶ M NaOH solution.

$$pOH = -log [1.85 \times 10-6M] = 5.733$$

 $pH = 14 - 5.73 = 8.267$

What is the [OH·] of an aqueous solution that has a pH of 5.0?

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35gNaOH}{2.000L} \times \frac{1molNaOH}{40.00gNaOH} = 0.44M NaOH$$

pOH = $-\log[0.44M] = 0.36$ pH = $14 - 0.36 = 13.64$ basic

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