

Percent Composition

Determine the percentage composition of sodium carbonate (Na_2CO_3)?

Molar Mass

$$\text{Na} = 2(23.00) = 46.0$$

$$\text{C} = 1(12.01) = 12.0$$

$$\text{O} = 3(16.00) = \underline{48.0}$$

$$\text{MM} = 106 \text{ g}$$

Percent Composition

$$\% \text{ Na} = \frac{46.0 \text{ g}}{106 \text{ g}} \times 100\% = 43.4 \%$$

$$\% \text{ C} = \frac{12.0 \text{ g}}{106 \text{ g}} \times 100\% = 11.3 \%$$

$$\% \text{ O} = \frac{48.0 \text{ g}}{106 \text{ g}} \times 100\% = 45.3 \%$$

Chemistry Solutions Manual Calculate The Percent Composition

**C. A. Trapp, Peter Atkins, Julio
dePaula, M. P. Cady, Carmen Giunta**



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. Regents Exams and Answers: Chemistry--Physical Setting Revised Edition

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