



"You can think about stoichiometry chemistry"
 • You can think of molecules as if as the ingredients to create the product (sauce)

Write a reaction below to represent the creation of a smoothie. Use Gr for grams, molecules, moles, and L for liters.



5 MORE STOICHIOMETRY SAMPLE PROBLEM

If you have 276 grams of water and an excess of dissolved sugar, how many grams of glucose can you make?

PH

<p>4 Acid Solution Contains fewer gas, sugar, fatty acid and protein acid</p>	<p>4 Neutral Solution Contains neither cooking oil, sugar solution</p>	<p>4 Basic Solution Contains fatty acids, cooking oils and sugar, bleach, soap, detergent</p>
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Acid solutions have a pH < 7
 Water does have a pH of 7
 Basic solutions have a pH > 7

Empirical formula mass = n = Molar Mass

Empirical formula CH_2O $\xrightarrow{\text{multiply subscripts by } n}$ Molecular formula $\text{C}_6\text{H}_{12}\text{O}_6$

If you have a molar mass of 360 grams, determine the molecular formula. Use the empirical formula from the previous page.

CHEMISTRY UNIT 9: STOICHIOMETRY DOODLE DIAGRAMS



Reactant 1 → Product
 Reactant 2 → Product

1. Balance atoms in the reactants.

2. Balance atoms in the products.

3. Check your work by counting atoms on both sides.

LIQUID REACTANT

1. Calculate the volume of the concentrated solution needed that contains that number of moles.

$$\left(\frac{100 \text{ mL}}{1 \text{ L}} \right) \left(\frac{10 \text{ L}}{10 \text{ moles}} \right) = 100 \text{ L of concentrated solution needed}$$

2. Add the volume of concentrated solution needed into a graduated cylinder or beaker and add water until the volume comes to the desired volume for the dilute solution.

FORMULA

Formula represents the atoms of each type of atom in a compound by molecular and empirical formula. An example is water. Water's atoms are H₂O and water's molecular formula is H₂O.

EMPIRICAL FORMULA

Empirical formula is used to define the smallest ratio of each atom in a compound.

MOLECULAR FORMULA

Molecular formula is used to define the actual number of each atom in a molecule.

GLUCOSE

$\text{C}_6\text{H}_{12}\text{O}_6$

Chemistry Semester 1 Unit 9 Stoichiometry Notes

Yicheng Fang



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