



Empirical formula CH_2O_2 $\xrightarrow{\text{multiply subscripts by } n}$ Molecular formula $\text{C}_3\text{H}_6\text{O}_6$

values in major veins of 100 g roots. Determine the molecular weight. Use the empirical formula from the previous step.

Chemistry Unit 9: Stoichiometry Doodle Diagrams



1. Write the balanced equation for the reaction.
 reactant 1 \rightarrow product 1
 reactant 2 \rightarrow product 2
 2. Decide which is the limiting reactant.
 3. Identify the product to reactant molar ratio from the chemical reaction information.
 4. Use the factor label method to calculate the amount of product that is possible to produce using the given reactant amount.
 5. Use the factor label method to calculate the amount of product that is possible to produce using the given reactant amount.

given:
 product 1
 (molar mass)
 $\frac{m}{1}$
 6. Use the factor label method to calculate the amount of product that is possible to produce using the given reactant amount.
 given:
 (molar mass)
 $\frac{m}{1}$

LIMITING REACTANT

Step 1: Calculate the volume of concentrated solution needed.

Calculate the volume of the concentrated solution needed that contains that number of moles:

$$\left(\frac{0.05 \text{ moles glucose}}{1} \right) \left(\frac{10 \text{ L}}{0.5 \text{ moles glucose}} \right) = 0.05 \text{ L of concentrated solution needed}$$

Step 2: Add the volume of concentrated solution needed into a graduated cylinder or beaker and add water until the volume comes to the desired volume for the whole solution.

The diagram illustrates the process of diluting a concentrated solution. It shows three beakers connected by arrows. The first beaker contains 0.05 L of 0.5 M glucose solution. An arrow points to the second beaker, which contains 10 L of water. Another arrow points to the third beaker, which contains 0.05 L of 0.05 M glucose solution, with the total volume marked as 10 L.

Abstract

Parasitic reproduction that occurs in small fragments of cotton is a consequence of inadequate seed separation. Parasitic seed samples in cotton bolls are common in the seed cotton.

is the number of the individual along the line of study (numbered 1 to 1000).

RESEARCH

and Karmali's used to argue that
an individual's health is not in the
control of the state, but is a result of
the state's failure to protect the health
of the individual.

1998

Questions



Chemistry Semester 1 Unit 9 Stoichiometry Notes

Yicheng Fang



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