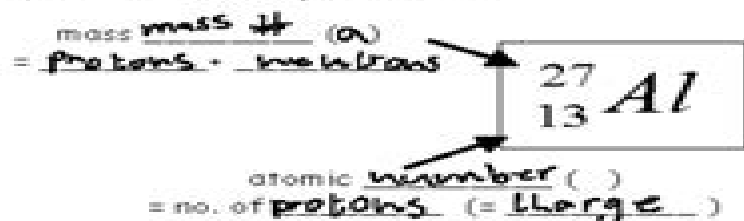


Nucleons:

- are particles occupying the nucleus
- consist of + charged protons and neutral charge neutrons
- have almost 2000 times the mass of electrons
- are made up of quark and leptons

Nuclear Notation:



$^{27}_{13}\text{Al}$ has 13 protons and 14 neutrons, for a total of 27 nucleons.

$^{28}_{13}\text{Al}$ has 13 protons and 15 neutrons, for a total of 28 nucleons.

isotopes: atoms of same element with different numbers of neutrons (different masses)

Forces Acting on Nucleons:

strong forces -

- forces of attraction between nucleons
- are independent of the charges of the nucleon
- are short range (exist only between closest neighbors)

electrical force -

- force of repulsion between + charged protons
- are long range

When are nuclei unstable?

- large nuclei ($Z > 82$) - electrical forces of repulsion are greater than strong forces of attraction
- wrong neutron : proton ratio

A radioactive isotope:

- has an unstable nucleus
- spontaneously emits a particle and charge into another element

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Table of Contents Chemistry Note Taking Guide Episode 1501 Answers

1. Understanding the eBook Chemistry Note Taking Guide Episode 1501 Answers
 - The Rise of Digital Reading Chemistry Note Taking Guide Episode 1501 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Note Taking Guide Episode 1501 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Note Taking Guide Episode 1501 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Note Taking Guide Episode 1501 Answers
 - Personalized Recommendations
 - Chemistry Note Taking Guide Episode 1501 Answers User Reviews and Ratings
 - Chemistry Note Taking Guide Episode 1501 Answers and Bestseller Lists
5. Accessing Chemistry Note Taking Guide Episode 1501 Answers Free and Paid eBooks

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 - Chemistry Note Taking Guide Episode 1501 Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Chemistry Note Taking Guide Episode 1501 Answers Compatibility with Devices
 - Chemistry Note Taking Guide Episode 1501 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemistry Note Taking Guide Episode 1501 Answers
 - Highlighting and Note-Taking Chemistry Note Taking Guide Episode 1501 Answers
 - Interactive Elements Chemistry Note Taking Guide Episode 1501 Answers
8. Staying Engaged with Chemistry Note Taking Guide Episode 1501 Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Note Taking Guide Episode 1501 Answers
9. Balancing eBooks and Physical Books Chemistry Note Taking Guide Episode 1501 Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Note Taking Guide Episode 1501 Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Note Taking Guide Episode 1501 Answers
- Setting Reading Goals Chemistry Note Taking Guide Episode 1501 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Note Taking Guide Episode 1501 Answers
- Fact-Checking eBook Content of Chemistry Note Taking Guide Episode 1501 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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