

Answer Question No. 10-12 (10 marks, 10% of total mark)

10. Calculate the pH of a 0.1 M solution of acetic acid (pKa = 4.75)

$$[H^+] = \sqrt{K_a \cdot C} = \sqrt{10^{-4.75} \cdot 0.1} = 1.32 \times 10^{-3} \text{ M}$$

$$pH = -\log(1.32 \times 10^{-3}) = 2.88$$

11. Calculate the pH of a 0.1 M solution of sodium acetate.

$$[OH^-] = \sqrt{K_b \cdot C} = \sqrt{10^{-9.25} \cdot 0.1} = 2.51 \times 10^{-5} \text{ M}$$

$$pOH = -\log(2.51 \times 10^{-5}) = 4.60$$

$$pH = 14 - pOH = 9.40$$

12. Calculate the pH of a 0.1 M solution of sodium acetate.

$$[OH^-] = \sqrt{K_b \cdot C} = \sqrt{10^{-9.25} \cdot 0.1} = 2.51 \times 10^{-5} \text{ M}$$

$$pOH = -\log(2.51 \times 10^{-5}) = 4.60$$

$$pH = 14 - pOH = 9.40$$

13. Calculate the pH of a 0.1 M solution of sodium acetate.

$$[OH^-] = \sqrt{K_b \cdot C} = \sqrt{10^{-9.25} \cdot 0.1} = 2.51 \times 10^{-5} \text{ M}$$

$$pOH = -\log(2.51 \times 10^{-5}) = 4.60$$

$$pH = 14 - pOH = 9.40$$

14. What is the conjugate base of acetic acid? CH_3COO^-

15. What is the conjugate acid of acetate? CH_3COOH

16. What is the conjugate base of HCl? Cl^-

17. Draw the structure of the acid and the following conjugate base:



18. Draw the structure of the acid and the following conjugate base:

- a. 1-methyl-2-propanol b. $CH_3CH_2CH_2O^-$
- c. 1-propanol d. $CH_3CH_2O^-$
- e. 2-propanol f. $CH_3CH(O^-)CH_3$
- g. 1-propanol h. $CH_3CH_2O^-$
- i. 1-propanol j. $CH_3CH_2O^-$
- k. 1-propanol l. $CH_3CH_2O^-$
- m. 1-propanol n. $CH_3CH_2O^-$
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