

Chemical Kinetics and Equilibrium

Sample Test Questions

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- Another name for the activated complex is _____.
A) energy barrier
B) collision group
C) rate limiter
D) transition state
- Activation energy is _____.
A) generally very high for a reaction that takes place rapidly
B) the heat released in a reaction
C) the energy given off when reactants collide
D) an energy barrier between reactants and products
- Why does a higher concentration make a reaction faster?
A) There are more collisions per second or the collisions are of greater energy.
B) Collisions occur with greater energy only.
C) There are more collisions per second and the collisions are of greater energy.
D) There are more collisions per second only.
- Why does a catalyst cause a reaction to proceed faster?
A) The activation energy is lowered only.
B) The collisions occur with greater energy only.
C) There are more collisions per second and the collisions are of greater energy.
D) There are more collisions per second only.
- What happens to a catalyst in a reaction?
A) It is unchanged.
B) It is incorporated into the products.
C) It is incorporated into the reactants.
D) It evaporates away.
- A catalyst works by _____.
A) changing the temperature of the reactants
B) shifting the equilibrium position toward the products
C) lowering the activation energy barrier
D) changing the particle size of the reactants
- The rate of a chemical reaction normally _____.
A) increases as reactant concentration increases
B) decreases as temperature increases
C) is slowed down by a catalyst
D) decreases as reactant concentration increases
- Which of the following substances act as catalysts in the body?
A) carbohydrates
B) lipids
C) nucleic acids
D) enzymes
- At equilibrium, what is the rate of production of reactants compared with the rate of production of products?
A) higher
B) lower
C) much higher
D) the same
- If a reaction is reversible, what are the relative amounts of reactant and product at the end of the reaction?
A) some product; some reactant
B) The relationship between reactants and products cannot be determined.
C) no product; all reactant
D) no reactant; all product
- If sulfur dioxide and oxygen can be made into sulfur trioxide, what is the reverse reaction?
A) $2\text{SO}_3 \rightarrow 2\text{SO}_2 + \text{O}_2$
B) $2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$
C) $\text{SO}_3 + \text{O}_2 \rightarrow \text{SO}_2$
D) $\text{SO}_2 + 2\text{SO}_3 \rightarrow 3\text{S} + 4\text{O}_2$
- Consider the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$. What is the effect of decreasing the volume on the contained gases?
A) The reaction shifts toward the product gas.
B) Ammonia is consumed in the reaction.
C) The pressure on the gases decreases momentarily.
D) The system reacts by increasing the number of gas molecules.

Chemistry Kinetics And Equilibrium Questions

Annelies Wilder-Smith



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