

CH301 Worksheet 1 Answer Key: High School Chemistry Questions.

For each problem, identify the kind of question BEFORE answering it.

1. A 25 kg ball has a kinetic energy of 225 J. What is the speed of the ball?
What kind of question is this? energy
3 m/s
2. How many molecules of water are there in 100 mL of water? (The density of water = 1.000 g/mL)
What kind of question is this? stoichiometry
 3.34×10^{24} molecules
3. Given that the atomic mass of carbon is given as 12.0107 g/mol, what are the approximate relative abundances of the two isotopes ^{12}C and ^{13}C ? *What kind of question is this? isotopes*
98.93% ^{12}C , 1.07% ^{13}C
4. In the electrolysis of $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$, how many atom of O in O_2 are made from 4 g of H_2O ?
What kind of question is this? stoichiometry
 1.34×10^{23} molecules
5. What are the most likely ionic forms of (a) sodium and (b) calcium?
What kind of question is this? periodic trends
(a) Na^+ ; (b) Ca^{2+}
6. What is the product of a reaction between metallic lithium and oxygen, O_2 ?
What kind of question is this? oxidation numbers
 Li_2O
7. Name the following species: (a) Cl^- , (b) NO_3^- , (c) C_3H_8 . *What kind of question is this? naming*
(a) chloride; (b) nitrate; (c) propane
8. The combustion of ethanol has the following equation: $\text{C}_2\text{H}_5\text{OH} + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O}$
What mass of CO_2 is formed from 1 kg of ethanol? *What kind of question is this? stoichiometry*
1.913 kg
9. Vitamin K contains 78.95% C, 3.95% H, and 21.05% O by mass. What is its empirical formula?
What kind of question is this? empirical formula calculation
 $\text{C}_{25}\text{H}_{30}\text{O}$
10. Given that vitamin K (see question 9) has molecular weight of 158.15 g/mol, what is its molecular formula? *What kind of question is this? molecular formula calculation*
 $\text{C}_{10}\text{H}_{14}\text{O}_2$

Chemistry If8766 Answer And Work

Marian DeWane, Heather Hattori



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Practice Makes Perfect Chemistry Review and Workbook, Second Edition Marian DeWane, Heather Hattori, 2018-12-28 The Winning Equation for Success in Chemistry is Practice Practice Practice This book will help you apply concepts and see how chemistry topics are interconnected Inside are numerous lessons to help you better understand the subject These lessons are accompanied by dozens of exercises to practice what you ve learned along with a complete answer key to check your work Throughout this book you will learn the terms to help you understand chemistry and you will expand your knowledge of the subject through hundreds of sample questions and their solutions With the lessons in this book you will find it easier than ever to grasp chemistry concepts And with a variety of exercises for practice you will gain confidence using your growing chemistry skills in your classwork and on exams **YOU LL BE ON YOUR WAY TO MASTERING THESE TOPICS AND MORE** Atomic structure The periodic table Chemical formulas Chemical reactions Mass and mole relationships Gas laws Solutions Acids and bases Thermochemistry A brand new chapter on the structure of molecules *U Can: Chemistry I For Dummies* John T. Moore, Chris Hren, Peter J. Mikulecky, 2015-07-21 Now you can score higher in chemistry Every high school requires a course in chemistry for graduation and many universities require the course for majors in medicine engineering biology and various other sciences U Can Chemistry I For Dummies offers all the how to content you need to enhance your classroom learning simplify complicated topics and deepen your understanding of often intimidating course material Plus you ll find easy to follow examples and hundreds of practice problems as well as access to 1 001 additional Chemistry I practice problems online As more and more students enroll in chemistry courses the need for a trusted and accessible resource to aid in study has never been greater That s where U Can Chemistry I For Dummies comes in If you re struggling in the classroom this hands on friendly guide makes it easy to conquer chemistry Simplifies basic chemistry principles Clearly explains the concepts of matter and energy atoms and molecules and acids and bases Helps you tackle problems you may face in your Chemistry I course Combines how to with try it to form one perfect resource for chemistry students If you re confused by chemistry and want to increase your chances of scoring your very best at exam time U Can Chemistry I For Dummies shows you that you can **Chemistry Work-book** Randal L. Henly, 1986 **Fundamentals of Chemistry** Fred H. Redmore, 1979 *Advanced Level Practical Work in Chemistry* John Parkinson, 1991 **Advanced Level Practical Work in Chemistry** John Parkinson, 1991

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