

Chemistry First Nine Weeks study guide for exam

1. What are SI base units for time, length, amount of substance, mass?
2. Express numbers in scientific notation given standard, and given scientific put in standard.
 $452.7 \text{ m} = ?$, $0.00077893 \text{ L} = ?$, $6.77 \times 10^4 \text{ mL} = ?$, $8.12 \times 10^{-6} \text{ km} = ?$
3. What are definitions of precision and accuracy and be able to apply them correctly to data given. See data tables in chapter 2 for examples.
4. Significant digits- how many in a given number, which types of zeroes are significant.
ex. 0.005603 m , 1241 m , 4.5800500 cm , 23.0 L
5. Correct rounding with significant digits. 4.5677 L to 3 sig. dig, 22.8655 to 4 sig.dig.
6. Adding, subtracting, multiplying, dividing correctly with sig. digits. (2 SEPARATE RULES)
 $34.41 \text{ m} + 11.8764 \text{ m} = ?$, $56.34 \text{ m} - 21.9 \text{ m} = ?$, $439 \text{ cm} \times 5.33 \text{ cm} = ?$, $189.22 \text{ g} / 76 \text{ m}$
7. Conversion of SI units to each other using dimensional analysis. You are responsible for the units you were given in class. Also make sure you know nano. Ex: $10^9 \text{ nm} = 1 \text{ m}$, very small unit, new branch of chemistry springing up around thee very tiny! You will hear more of this in future.
8. Density problems. Find the third term if given the other 2. Know the formula; it won't be given. Find density of a rock with mass of 35.8 g and volume of 8.1 mL . Find the volume of a substance with mass of 163 grams and density of 5.77 g/mL .
9. Dimensional analysis problems. See review problem sets for examples.
10. Characteristics of the 3 basic states of matter as they pertain to their container, volume, shape, etc.
11. Recognition of physical properties and chemical properties and/or changes. Copper carbonate which is green in color, when heated, turns black and loses mass. Grinding charcoal into a fine powder is an example of which?
12. What principles and means are used to separate mixtures?
13. Distinguish between homogeneous and heterogeneous mixtures. Well mixed koolaid, tossed salad, unsweetened iced tea, orange juice with pulp, solution.
14. Law of Conservation of Mass. State it and explain how you can use it to solve basic problems. Also % mass of an element in a compound. See practice problem sets.
15. State 4-5 parts of Dalton's atomic theory. Whose idea had he revived? Which part(s) do we not agree with completely anymore?
16. Three subatomic particles in an atom and where they are located. Be able to find all three of them if given atomic mass (and PT). Know how to write both hyphen notation and nuclear notation.
16. What were the characteristics of the nucleus that Rutherford discovered? What experiment did he discover them in?

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L Reisser



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problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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