

# Chemistry 101 Final Exam Study Guide- Spring 2017

## Acids and Bases

Acid: substance that provides  $H^+$  ions in water

Base: substance that provides  $OH^-$  ions in water

## Covalent Bonds

- Review: ionic bonds are formed by the positive charge of the cation being attracted to the negative charge of the anion
- Covalent bonds: bonds formed by sharing electrons between the atoms (not ions)
  - 1 electron from each atom (2 atoms per bond)
  - Electrons available for bonding are the un-paired electrons in orbitals (orbitals with only 1 electron in them)
- Molecule
  - A group of atoms held together by covalent bonds
- Atoms will bond with enough other atoms to complete the octet
  - But they can only form as many bonds as unpaired valence electrons
    - Hydrogen has 1 valence electron, so it can only form 1 bond, and will not achieve an octet
    - Boron has 3 valence electrons, so can only form 3 bonds, and will not achieve the octet (only 6 electrons)
  - There are always exceptions to the rules
    - Sulfur and phosphorus can extend the octet

## Multiple Covalent Bonds between 2 atoms

- if the octet is not met, additional bonds may be formed
  - $CO_2$  (carbon dioxide)
    - C has 4 valence electrons, so 4 bonds possible
    - O has 6 valence electrons, 2 are unpaired, so 2 bonds are possible

## PRACTICE

### IONS:

- Sodium
  - Symbol: Na
  - Valence: 1
  - Electron dot:  $Na^+$
  - Ionic dot:  $Na^+$  (loses its electron to get a 1+ charge)
- Magnesium
  - Symbol: Mg
  - 2 valence  $e^-$
  - Loses 2 electrons to get 2+ charge
- Oxygen
  - Symbol: O
  - Valence: 6
  - Will gain 2 electrons for a 2- charge

Cannot have a covalent bond with ions, just atoms

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**Chandresh Agrawal, nandini books**



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