# Chemistry 338 2<sup>nd</sup> Semester Final Exam Review Sheet

#### Unit 1 - Matter & Measurement - Chapters 1 & 2

 A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

% error =  $\frac{experimentat-accepted}{accepted} \times 100$  % error =  $\frac{4.2-5.0}{5.0} \times 100 = -16\%$  error

#### Unit 2 - Periodic Table - Chapter 5

2.

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass#	Charge
Fungsten ion	194 W+2	7-4	74	110	72	184	+2

 Using the following atomic radius values, predict the atomic radius of potassium (K): Li = 145 pm, Na = 180 pm, K = 7, Rb = 235 pm K = 220 pm

#### Unit 3 - Reactions/Compounds - Chapters 7 & 8

- 4. Classify the compound as ionic or covalent. Name the compound
  - Fe<sub>2</sub>O<sub>3</sub> iron (III) oxide (ionic roman numerals for transition metals, change ending to -ide)
  - b. Cl<sub>2</sub>O<sub>8</sub> dichlorine octaoxide (covalent use prefixes, change ending to -ide)
- 5. Classify the compound as ionic or covalent. Write the formula
  - a. Trisulfur Dinitride S<sub>3</sub>N<sub>2</sub> (covalent)
  - b. Zinc Hydroxide Zn(OH)2 (ionic look up ion charges and criss-cross) Zn<sup>-2</sup> OH<sup>-3</sup>
  - c. Chromium (III) nitrate Cr(NO<sub>3</sub>)<sub>3</sub> (ionic criss-cross) Cr<sup>23</sup> NO<sub>3</sub><sup>-1</sup>
- 6. Write the equation, balance it, and classify the type of reaction:

Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate

> 1 FeSO<sub>4</sub> + 2 NaOH → 1 Fe(OH)<sub>2</sub> + 1 Na<sub>2</sub>SO<sub>4</sub> double displacement (two ionic compounds reacting)

7. Balance the equation and classify the type of reaction:

1 Cu + 2 AgNO<sub>3</sub> → 1 Cu(NO<sub>3</sub>)<sub>2</sub> + 2 Ag Single displacement (element and a compound reacting)

#### Unit 4 - Moles - Chapter 3

8. Calculate the molar mass of Fe<sub>2</sub>(SO<sub>4</sub>)<sub>5</sub>.

Fe (2)(55.9 g/mol) = 111.8 g/mol

S (3)(32.1 g/mol) = 96.3 g/mol

+ O (12)(16.0 g/mol) = +192 g/mol

= 400.1 g/mol

# **Chemistry Final Exam Review Semester 2 Answers**

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