Chemistry Final Exam Review Questions

Useful Values and Equations

Some of the following values and equations may be useful for answering some of the questions and problems.

9P = 1.8(9C) + 32speed of light = 3.00×10^8 m/s $c = \lambda w$

R = 0.082057 L-atm / deg-mole Planck's constant = 6.63 × 10-34 J·s Fig. 1800

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- One of the following does not describe gases. Which one is it?
 - a. Expands infinitely.

Easily compressed.

b. Rigid shape.

- d. Relatively low densities.
- The following properties describe the element, mercury (Hg). Which one is a chemical property?
 - a. Its density is 5.43 g/cm³.
 - b. It is a silvery-white metal.

- c. It is a liquid at 298 K.
- d. Mercury(II) sulfide forms when it is exposed to sulfide ions.
- All of the following are properties of SO₂. Which one is a physical property?
 - It reacts with oxygen to form SO₃².

- e. It is a strong reducing agent.
- mospheric pressure.
- It condenses to a colorless liquid at 14°F and at It reacts with water and oxygen to form an acid...
- 4. Which of the following numbers has 4 significant figures?

b. 0.0430

c. 0.0431

- 0.43980
- 5. How many significant figures should there be in the answer to the following math equation? 2.834 + 5.71520 + 2.12 + 178.1 + 250.2619
 - a. 3

e. 5

- 6. How many significant figures should there be in the answer to the following math equation? $23.413 \text{ g} + (2.15 \text{ cm} \times 1.1 \text{ cm} \times 3.73 \text{ cm})$

b. 3

- 7. At 25°C, one cm3 of iron has a mass of 7.62 grams. What volume of iron is occupied by 125 grams of iron?
 - a. 0.82 cm³

- b. 8.2 cm³
- c. 0.164 cm³

- d. 16.4 cm³
- 8. A cube of lead has a volume of 6.61 cm3. Its mass is 75.0 grams. What is its density?
 - a. 3.85 g/cm³
- b. 7.41 g/cm³
- c. 11.3 g/cm³
- d. 22.6 g/cm³

Chemistry Final Exam Review Questions

Roman Wölfel

Chemistry Final Exam Review Questions:

This book delves into Chemistry Final Exam Review Questions. Chemistry Final Exam Review Questions is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemistry Final Exam Review Questions, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 1: Introduction to Chemistry Final Exam Review Questions
 - Chapter 2: Essential Elements of Chemistry Final Exam Review Questions
 - Chapter 3: Chemistry Final Exam Review Questions in Everyday Life
 - Chapter 4: Chemistry Final Exam Review Questions in Specific Contexts
 - ∘ Chapter 5: Conclusion
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- 6. In chapter 5, this book will draw a conclusion about Chemistry Final Exam Review Questions. The final chapter will summarize the key points that have been discussed throughout the book.

 This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Chemistry Final Exam Review Questions.

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