

# EXPERIMENT NO. 1

**Aim :** Preparation of :

- (I) a true solution of common salt, sugar and alum.
- (II) a suspension of soil, chalk powder and fine sand in water.
- (III) a colloidal solution of starch in water and egg albumin/milk in water and distinction between these on the basis of :
  - transparency
  - filtration criterion
  - stability

## Concept & Theory :

A true solution is a homogeneous mixture of two or more substances (or of a solute and a solvent). Of the substances, the one which is present in the largest amount is called the solvent. The other substance present in a smaller quantity is called the solute.

**Solubility :** The maximum amount of a solute dissolved in 100 g of a solvent under a particular condition of temperature and pressure is the solubility of the solute in that solvent under the specified conditions.

**Saturation :** When no more solute can be dissolved in a solvent at the given condition of temperature and pressure, the solution is said to be saturated or to have reached the saturation point.

In a solution, the solute may be completely soluble, partially solution, or completely insoluble. Depending on the dispersion of a solute in a solvent, the resultant solution may be homogeneous (solution) or heterogeneous material (a suspension or a colloid).

## Objectives :

To perform a set of experiments in order to understand the characteristics (transparency, filterability, stability) of and the differences among a true solution, suspension, and a colloid.

## Before We Start–Take Precautions :

Handle acids and strong agents carefully.

Use a glass rod dipped in a funnel to direct the path of content when filtering a solution through filter paper.

Do not shake test tubes at all during the stability test period.

## (I) Preparation of a true solution of common salt, sugar and alum

**Materials** (Assemble the following on your work-bench before you start) :

Apparatus	Chemicals/Agents
Beakers x 3	Distilled water
Glass rod x 1	10 g common salt (powder)
	10 g sugar (powder)
	10 g alum (powder)

## Procedure :

1. In three beakers, add 90 ml water each, and label the beakers as A, B and C.
2. Add 10 g of finely powdered common salt, sugar and alum in beakers A, B and C respectively.
3. Stir the contents of each beaker using a glass rod and set aside until further use.
4. This solution is a true solution.

## (II) Preparation of a suspension of soil, chalk powder and fine sand in water

**Materials** (Assemble the following on your work-bench before you start) :

Apparatus	Chemicals/Agents
Beakers x 3	Distilled water
Glass rod x 1	10 g fine soil
	10 g chalk powder
	10 g fine sand

# Chemistry Experiment 1 Class 9 Laboratory Manual

**Daniela Niemeyer**



## Chemistry Experiment 1 Class 9 Laboratory Manual:

*Core Science Lab Manual with Practical Skills for Class IX* V. K. Sally, Chhaya Srivastava, Goyal Brothers Prakashan, 2019-01-01 Goyal Brothers Prakashan EduGorilla's CBSE Class 11th Chemistry Lab Manual | 2024 Edition | A Well Illustrated, Complete Lab Activity book with Separate FAQs for Viva Voce Examination , Need an informative and well illustrated Lab Manual CBSE Class 11th Chemistry Lab Manual is here for you The Lab Manual provides comprehensive steps for guiding students through each experiment Rigorously researched content prepared by a team of educators writers editors and proofreaders CBSE Class XI Chemistry Lab Manual has properly labeled high resolution diagrams and graphs A separate section on Viva Questions has been included to aid students in their Viva examination The Lab Manual explains the complex topics through detailed illustrations and lucid language making them simple to grasp Worksheets have been provided in CBSE Class 11th Chemistry Lab Manual for doing rough work **Chemistry Lab Manual** Neena Sinha, R Rangarajan, R P Manchanda, R K Gupta, Rajesh Kumar, Lab Manual **EduGorilla's CBSE Class 12th Chemistry Lab Manual | 2024 Edition | A Well Illustrated** EduGorilla Prep Experts, **Laboratory Manual for Principles of General Chemistry** J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section *Chemistry Lab Manual Class XI | follows the latest CBSE syllabus and other State Board following the CBSE Curriculam.* Mr. Rohit Manglik, Mr. Kaushalesh Dwivedi, 2022-08-04 With the NEP and expansion of research and knowledge has changed the face of education to a great extent In the Modern times education is not just constricted to the lecture method but also includes a practical knowledge of certain subjects This way of education helps a student to grasp the basic concepts and principles Thus trying to break the stereotype that subjects like Physics Chemistry and Biology means

studying lengthy formulas complex structures and handling complicated instruments we are trying to make education easy fun and enjoyable

**Hard Bound Lab Manual Chemistry** Neena Sinha, R Rangarajan, R P Manchanda, R K Gupta, Rajesh Kumar, Lab Manuals *Laboratory Manual for Principles of General Chemistry* Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

Lab Manual for General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, 2009-08-21 Teaching all of the necessary concepts within the constraints of a one term chemistry course can be challenging Authors Denise Guinn and Rebecca Brewer have drawn on their 14 years of experience with the one term course to write a textbook that incorporates biochemistry and organic chemistry throughout each chapter emphasizes cases related to allied health and provides students with the practical quantitative skills they will need in their professional lives Essentials of General Organic and Biochemistry captures student interest from day one with a focus on attention getting applications relevant to health care professionals and as much pertinent chemistry as is reasonably possible in a one term course Students value their experience with chemistry getting a true sense of just how relevant it is to their chosen profession To browse a sample chapter view sample ChemCasts and more visit [www.whfreeman.com/gob](http://www.whfreeman.com/gob)

*Practical/Laboratory Manual Chemistry Class XII based on NCERT guidelines by Dr. S. C. Rastogi, Er. Meera Goyal* Dr. S. C. Rastogi, Er. Meera Goyal, 2020-06-22

**A Surface Chemistry**

- 1 To prepare colloidal solution sol of starch
- 2 To prepare a colloidal solution of egg albumin
- 3 To prepare colloidal solution of gum
- 4 To prepare colloidal solution of aluminium hydroxide  $\text{Al}(\text{OH})_3$
- 5 To prepare colloidal solution of ferric hydroxide  $\text{Fe}(\text{OH})_3$
- 6 To prepare colloidal solution of arsenious sulphide  $\text{As}_2\text{S}_3$
- 7 To purify a freshly prepared sol by dialysis
- 8 To compare the effectiveness of different common oils Castor oil cotton seed oil coconut oil kerosene oil mustard oil in forming emulsions

**Viva Voce B Chemical Kinetics**

- 1 To study the effect of concentration on the rate of reaction between sodium thiosulphate and hydrochloric acid
- 2 To study the effect of temperature on the rate of reaction between sodium thiosulphate and hydrochloric acid
- 3 To study the rate of reaction of iodide ions with hydrogen peroxide at different concentrations of iodide ions
- 4 To study the rate of reaction between potassium iodate  $\text{KIO}_3$  and sodium sulphite  $\text{Na}_2\text{SO}_3$  using starch solution as indicator

**Viva Voce C Thermochemistry**

- 1 Determine the enthalpy of dissolution of copper sulphate  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  in water at Room temperature
- 2 To determine the enthalpy of neutralization of the reaction between  $\text{HCl}$  and  $\text{NaOH}$
- 3 To determine enthalpy change during the interaction between acetone and chloroform

**Viva Voce D Electrochemistry**

- 1 To study the variation of cell potential in  $\text{Zn}|\text{Zn}^{2+}||\text{Cu}^{2+}|\text{Cu}$  with change in concentration of electrolytes  $\text{CuSO}_4$  or  $\text{ZnSO}_4$  at room temperature

**Viva Voce E Chromatography**

- 1 To separate the coloured components pigment present in the given extract of leaves and flowers by ascending paper chromatography and find their  $R_f$  values
- 2 To separate the

coloured components present in the mixture of red and blue inks by ascending paper chromatography and find their  $R_f$  values 3 To separate  $Co^{2+}$  and  $Ni^{2+}$  ions present in the given mixture by using ascending paper chromatography and determine their  $R_f$  values Viva Voce F Preparation of Inorganic Compounds 1 Preparation of double salt of ferrous ammonium sulphate Mohr's salt from ferrous sulphate and ammonium sulphate 2 To prepare a pure sample of potash alum fitkari 3 Preparation of crystals of potassium ferric oxalate or potassium trioxalato ferrate III Viva Voce G Preparation of Organic Compounds 1 Preparation of iodoform from ethyl alcohol or acetone 2 Preparation of acetanilide in laboratory 3 Preparation of  $\beta$ -Naphthol aniline dye 4 To prepare a pure sample of dibenzalacetone 5 To prepare a pure sample of  $p$ -nitro acetanilide Viva Voce H Tests for the Functional Groups Present in Organic Compounds Viva Voce I Study of Carbohydrates Fats and Proteins 1 To study simple reactions of carbohydrate 2 To study simple reactions of fats 3 To study simple reactions of proteins 4 To investigate presence of carbohydrates fats and proteins in food stuffs Viva Voce J Volumetric Analysis 1 To prepare 250 ml of  $M/10$  solution of oxalic acid 2 To prepare 250 ml of  $M/10$  solution of ferrous ammonium sulphate 3 Prepare  $M/20$  solution of oxalic acid with its help find out the molarity and strength of the given solution of potassium permanganate 4 Prepare  $M/20$  solution of Mohr's salt using this solution determine the molarity and strength of potassium permanganate solution Viva Voce K Qualitative Analysis Viva Voce INVESTIGATORY PROJECTS 1 To study the presence of oxalate ions in guava fruit at different stages of ripening 2 To study the quantity of caseine present in different samples of milk 3 Preparation of soyabean milk and its comparison with natural milk with respect to curd formation effect of temperature etc 4 To study the effect of potassium bisulphite as food preservative at various concentrations 5 To study the digestion of starch by salivary amylase and the effect of pH and temperature on it 6 To study and compare the rate of fermentation of the following materials wheat flour gram flour potato juice and carrot juice 7 To extract essential oils present in saunf aniseed ajwain corum illaichi cardamom 8 To detect the presence of adulteration in fat oil and butter 9 To investigate the presence of  $NO_2$  in brinjal

**Environmental Chemistry in the Lab** Ruth Ann Murphy, 2022-08-31 Environmental Chemistry in the Lab presents a comprehensive approach to modern environmental chemistry laboratory instruction together with a complete experimental experience The laboratory experiments have an introduction for the students to read a pre lab for them to complete before coming to the lab a data sheet to complete during the lab and a post lab which would give them an opportunity to reinforce their understanding of the experiment completed Instructor resources include a list of all equipment and supplies needed for 24 students a lab preparation guide an answer key to all pre lab and post lab questions sample data for remote learners and a suggested rubric for grading the labs Additional features include Tested laboratory exercises with instructor resources for environmental science students Environmental calculations industrial regulation and environmental stewardship Classroom and remote exercises An excellent user friendly and thought provoking presentation which will appeal to students with little or no science background A qualitative approach to the chemistry behind many of our environmental issues today *Core*

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the 3d Ed. of "A Course in General Chemistry" William McPherson, William Edwards Henderson, William Lloyd Evans, 1928  
*The Journal of Industrial and Engineering Chemistry* ,1921     **Prentice Hall Physical Science Concepts in Action**  
**Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action  
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 content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and  
 into the world around them Now includes even more technology tools and activities to support differentiated instruction  
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 experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both  
 written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked  
 to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations  
**Journal of Industrial and Engineering Chemistry** ,1921     **Resources in Education** ,1987-07     *Journal of*  
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