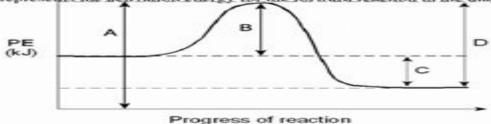
PART I Total Value: 50%

Instructions: Shade the letter of the correct answer on the computer scorable answer sheet provided.

Why does increasing temperature increase reaction rate?

(A) decreases activation energy

- (B) decreases the percentage of effective collisions
- (C) increases activation energy
- (D) increases the percentage of effective collisions
- 2. Which represents the activation energy for the forward reaction in the diagram below?



- (A) A (B) B
 - (C) C (D) D
- 3. Which type of change is best for monitoring the reaction rate of the reaction below?

$$Mg(OH)_2(s) + 2 HCl(aq) \rightarrow 2 H_2O(\ell) + MgCl_2(aq)$$

- (A) colour (B) mass
- (C) pH (D) volume
- 4. Which describes an activated complex?
 - (A) has a lower potential energy than products (B) has a lower potential energy than reactants
- (C) stable species that forms upon collision of reactants
 (D) unstable species that forms upon collision of reactants
- 5. According to the collision theory, why does the reaction below occur in more than one step?

$$C_5H_{12}(g) + 8 O_2(g) \rightarrow 5 CO_2(g) + 6 H_2O(g)$$

- (A) low C₃H₁₂(g) concentration
 - (B) low probability of a nine-particle collision
 - (C) particles collide with insufficient kinetic energy
 - (D) temperature is very high

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