# Chapter 10 Outline Chemical Quantities

<ul> <li>Section</li> </ul>	10.1 – The Mole: A Mea	surement of Matter			
<ul> <li>You ofte</li> </ul>	n measure the amount o	of something by	Practice Problems		
	, by, or by _				
A (mol) of a substance is			<ul> <li>How many moles are in 4.65 x 10<sup>34</sup> molecules of NO<sub>2</sub>?</li> </ul>		
represer	ntative particles of that s	ubstance.			
6.02 x 10	0 <sup>33</sup> is called	number.		J-2-2 G	
• 1 mole =			<ul> <li>How many atoms are</li> </ul>	in 4.33 mot magn	esium sulfate?
• A		refers to the species			
	in a substance: usually _			Cara Cara Cara Cara	2000000000000
			• The		
<ul> <li>Element</li> </ul>	Elements normally exist as, but 7 elements		grams is the mass of a		
	molecu		The mass of a	_ of an element i	s the
		5" 14			
Sample	Problem		To calculate the		
	How many moles is 2.80 x 19 <sup>24</sup> atoms of silicon?		the number of		The second secon
			one mole of the comp	ound. Then	the masses of
			the elements in the compound.		
Practice Problems			Sample Problem		
	How many moles is 2.17 x 10 <sup>23</sup> representative particles		<ul> <li>What is the molar mass of PCI<sub>3</sub>?</li> </ul>		
of bromine?					
OI DIOIN	iner		<ul> <li>Practice Problems</li> </ul>		
			<ul> <li>What is the molar mass of sodium hydrogen carbonate?</li> </ul>		
<ul> <li>How ma</li> </ul>	How many molecules are in 2.12 mol of propane? (m/c molecules)		. Miles is the sense of or	alekum miteatu 3	
w molecu			<ul> <li>What is the mass of calcium nitrate?</li> </ul>		

Sample Problem

How many atoms are in 1.14 mol SO<sub>3</sub>?

# **Chemistry Chemical Quantities Study Guide**

Barron's Educational Series, Mark Kernion, Joseph A. Mascetta

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verbesserung der ausdauerleistung erzielt

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