

Chapter 8

Redox Reaction

But, when conc. H_2SO_4 is added to an inorganic mixture containing chloride, a pungent smelling gas (HCl) is evolved. HCl , being a weak reducing agent, cannot reduce H_2SO_4 to SO_2 .



Question 8.13:

Identify the substance oxidised reduced, oxidising agent and reducing agent for each of the following reactions:

- (a) $2\text{AgBr}(s) + \text{C}_6\text{H}_6\text{O}_2(aq) \rightarrow 2\text{Ag}(s) + 2\text{HBr}(aq) + \text{C}_6\text{H}_4\text{O}_2(aq)$
 (b) $\text{HCHO}(l) + 2[\text{Ag}(\text{NH}_3)_2]^+(aq) + 3\text{OH}^-(aq) \rightarrow 2\text{Ag}(s) + \text{HCOO}^-(aq) + 4\text{NH}_3(aq) + 2\text{H}_2\text{O}(l)$
 (c) $\text{HCHO}(l) + 2\text{Cu}^{2+}(aq) + 5\text{OH}^-(aq) \rightarrow \text{Cu}_2\text{O}(s) + \text{HCOO}^-(aq) + 3\text{H}_2\text{O}(l)$
 (d) $\text{N}_2\text{H}_4(l) + 2\text{H}_2\text{O}_2(l) \rightarrow \text{N}_2(g) + 4\text{H}_2\text{O}(l)$
 (e) $\text{Pb}(s) + \text{PbO}_2(s) + 2\text{H}_2\text{SO}_4(aq) \rightarrow 2\text{PbSO}_4(s) + 2\text{H}_2\text{O}(l)$

Answer 8.13:

- (a) Oxidised substance $\rightarrow \text{C}_6\text{H}_6\text{O}_2$
 Reduced substance $\rightarrow \text{AgBr}$
 Oxidising agent $\rightarrow \text{AgBr}$
 Reducing agent $\rightarrow \text{C}_6\text{H}_6\text{O}_2$
 (b) Oxidised substance $\rightarrow \text{HCHO}$
 Reduced substance $\rightarrow [\text{Ag}(\text{NH}_3)_2]^+$
 Oxidising agent $\rightarrow [\text{Ag}(\text{NH}_3)_2]^+$
 Reducing agent $\rightarrow \text{HCHO}$
 (c) Oxidised substance $\rightarrow \text{HCHO}$
 Reduced substance $\rightarrow \text{Cu}^{2+}$
 Oxidising agent $\rightarrow \text{Cu}^{2+}$
 Reducing agent $\rightarrow \text{HCHO}$
 (d) Oxidised substance $\rightarrow \text{N}_2\text{H}_4$
 Reduced substance $\rightarrow \text{H}_2\text{O}_2$
 Oxidising agent $\rightarrow \text{H}_2\text{O}_2$
 Reducing agent $\rightarrow \text{N}_2\text{H}_4$

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Thomas Griffiths



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