

General Chemistry II Chapter 17 Study Guide

Organic acids- carboxylic acids (R-COOH):

- Carboxylic acids (R-COOH): a weak organic acid; contains the carboxyl group (-COOH); never fully dissociates in water; an equilibrium process; ex. Formic acid, acetic acid, butyric acid, benzoic acid
 - R= rest of the molecule
- Strength of carboxylic acids: depends on the R group; more electronegative Cl in R group shifts more electron density towards the R group and helps stabilize the structure after the H^+ ion is removed; thus, chloroacetic acid is the stronger acid in comparison to acetic acid (CH_3COOH)

Organic bases- amines:

- Amines: organic bases; contains an N atom with a lone pair (ex. $-NH$, $-NH_2$); ex. Methyl amine, dimethyl amine, trimethyl amine; are weak bases:
 - Only N with lone pair can accept a proton
 - Equilibrium dissociation process
 - NH_4^+ is an acid and can lose one H^+ to be neutral

Organic acids and bases (things to keep in mind):

- Acid: H is the only acidic proton; ex. CH_3COOH becomes CH_3COO^-
- Amine: the proton H^+ is bonded to N via the lone pair of N; ex. CH_3NH_2 becomes $CH_3NH_3^+$
- How about carbon?
 - Protons on C are not acidic; C cannot accept extra protons

Autoionization of water:

- Autoionization: in pure water, one water molecule can donate a proton to another water molecule



$$\text{At } 25^\circ\text{C, } [H_3O^+] = [OH^-] = 1.0 \times 10^{-7} \text{ M}$$

$$K_w = [H_3O^+] [OH^-] = 1.0 \times 10^{-14}$$

K_w : ion-product constant

Chemistry Chapter 17 Study Guide Answers

S Ben Porath



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Table of Contents Chemistry Chapter 17 Study Guide Answers

1. Understanding the eBook Chemistry Chapter 17 Study Guide Answers
 - The Rise of Digital Reading Chemistry Chapter 17 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 17 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Chapter 17 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 17 Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chapter 17 Study Guide Answers User Reviews and Ratings
 - Chemistry Chapter 17 Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Chapter 17 Study Guide Answers Free and Paid eBooks
 - Chemistry Chapter 17 Study Guide Answers Public Domain eBooks
 - Chemistry Chapter 17 Study Guide Answers eBook Subscription Services
 - Chemistry Chapter 17 Study Guide Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 17 Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 17 Study Guide Answers Compatibility with Devices
 - Chemistry Chapter 17 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 17 Study Guide Answers
 - Highlighting and Note-Taking Chemistry Chapter 17 Study Guide Answers
 - Interactive Elements Chemistry Chapter 17 Study Guide Answers

8. Staying Engaged with Chemistry Chapter 17 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 17 Study Guide Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 17 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 17 Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 17 Study Guide Answers
 - Setting Reading Goals Chemistry Chapter 17 Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 17 Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chapter 17 Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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