General Chemistry II Chapter 17 Study Guide

Organic acids- carboxylic acids (R-COOH):

- Carboxylic acids (R-COOH): a weak organic acid; contains the carboxyl group (-COOH); never fully dissociates in water, an equilibrium process; ex. Formic acid, acetic acid, butyric acid, benzoic acid
 - R= rest of the molecule
- Strength of carboxylic acids: depends on the R group; more electronegative CI in R group shifts more electron density towards the R group and helps stabilize the structure after the H* ion is removed; thus, chloroacetic acid is the stronger acid in comparison to acetic acid (CH-COOH)

Organic bases- amines:

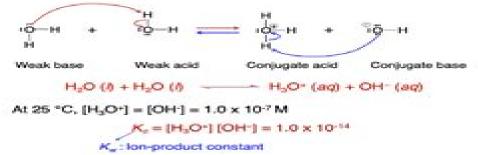
- Amines: organic bases; contains an N atom with a lone pair (ex. -NH, -NH₂); ex. Methyl amine, dimethyl amine, trimethyl amine; are weak bases:
 - Only N with lone pair can accept a proton
 - Equilibrium dissociation process
 - NH₄^{*} is an acid and can lose one H^{*} to be neutral

Organic acids and bases (things to keep in mind):

- Acid: H is the only acidic proton; ex. CH₂COOH becomes CH₂COO
- Amine: the proton H* is bonded to N via the lone pair of N; ex. CH₃NH₂ becomes CH₃NH₃*
- How about earbon?
 - Protons on C are not acidic; C cannot accept extra protons

Autoionization of water:

 Autoionization: in pure water, one water molecule can donate a proton to another water molecule



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S Ben Porath

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Table of Contents Chemistry Chapter 17 Study Guide Answers

- 1. Understanding the eBook Chemistry Chapter 17 Study Guide Answers
 - The Rise of Digital Reading Chemistry Chapter 17 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry Chapter 17 Study Guide Answers
 - $\circ \ Exploring \ Different \ Genres$
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 17 Study Guide Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry Chapter 17 Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chapter 17 Study Guide Answers User Reviews and Ratings
 - Chemistry Chapter 17 Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Chapter 17 Study Guide Answers Free and Paid eBooks
 - o Chemistry Chapter 17 Study Guide Answers Public Domain eBooks
 - Chemistry Chapter 17 Study Guide Answers eBook Subscription Services
 - Chemistry Chapter 17 Study Guide Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 17 Study Guide Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chemistry Chapter 17 Study Guide Answers Compatibility with Devices
 - Chemistry Chapter 17 Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Chemistry Chapter 17 Study Guide Answers
 - Highlighting and Note-Taking Chemistry Chapter 17 Study Guide Answers
 - Interactive Elements Chemistry Chapter 17 Study Guide Answers

- 8. Staying Engaged with Chemistry Chapter 17 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 17 Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 17 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 17 Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 17 Study Guide Answers
 - Setting Reading Goals Chemistry Chapter 17 Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 17 Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chapter 17 Study Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - $\circ \ \ Integration \ of \ Multimedia \ Elements$
 - Interactive and Gamified eBooks

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