

Preparing for the Organic Chemistry Exam?

1. Basic Concepts

Atomic Structure: Atoms are made of three sub-particles: Electrons, Protons, and Neutrons.

Electron: Negatively charged particle, mass 9.1×10^{-31} kg.

Proton: Positively charged particle, mass 1.67×10^{-27} kg.

Neutron: Neutral particle, mass 1.67×10^{-27} kg.

Atomic Number (Z): Number of protons in the nucleus.

Mass Number (A): Sum of protons and neutrons in the nucleus.

Isotopes: Atoms of the same element with same atomic number but different mass numbers.

2. Atomic Structure of Matter

Atomic Model: The atom consists of a central nucleus surrounded by electrons.

Electron Configuration: Electrons are arranged in shells or orbitals.

Bohr's Model: Electrons move in fixed orbits around the nucleus.

Spectral Lines: Discrete lines of light emitted or absorbed by atoms.

3. Periodic Table

Periodic Law: Properties of elements are a periodic function of their atomic number.

Groups: Vertical columns of elements with similar properties.

Periods: Horizontal rows of elements.

Metals: Elements on the left side of the periodic table.

Non-metals: Elements on the right side of the periodic table.

Alloids: Solid solutions of two or more metals.

4. Chemical Bonding

Electron Dot Diagram: Diagram showing the arrangement of valence electrons in atoms.

Valency: The combining capacity of an element.

Chemical Formula: Symbolic representation of a chemical compound.

Molecular Formula: Formula showing the actual number of atoms of each element in a molecule.

Structural Formula: Diagram showing the arrangement of atoms in a molecule.

5. Acids and Bases

Acids: Substances that release H^+ ions in aqueous solution.

Bases: Substances that release OH^- ions in aqueous solution.

pH Scale: Scale from 0 to 14 used to measure the acidity or basicity of a solution.

Indicators: Substances that change color in acidic or basic solutions.

Neutralization: Reaction between an acid and a base to form a salt and water.

6. Metals and Non-metals

Metals: Elements that are lustrous, malleable, and good conductors of heat and electricity.

Non-metals: Elements that are brittle, poor conductors, and have low melting and boiling points.

Alloys: Homogeneous mixtures of two or more metals.

Corrosion: The process of deterioration of metals due to reaction with the environment.

Prevention of Corrosion: Methods to protect metals from corrosion.

7. Carbon and its Compounds

Carbon: Element with atomic number 6, found in various forms like diamond, graphite, and fullerenes.

Carbon Compounds: Compounds containing carbon, such as hydrocarbons, alcohols, and acids.

Hydrocarbons: Compounds containing only carbon and hydrogen.

Alcohols: Compounds containing a hydroxyl group ($-OH$).

Acids: Compounds that release H^+ ions in aqueous solution.

8. Biomolecules

Carbohydrates: Organic compounds containing carbon, hydrogen, and oxygen.

Proteins: Large biomolecules made of amino acids.

Nucleic Acids: Biomolecules that store and transmit genetic information.

Enzymes: Biological catalysts that speed up chemical reactions.

Vitamins: Organic compounds that are essential for the growth and health of the body.

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