

**Section 11.4 Empirical and Molecular Formulas**

In your textbook, read about percent composition.

Answer the following questions.

1. What is the percent composition of a compound?

The percent of the total mass of the compound that is due to that component.

2. Describe how to find the percent composition of a compound if you know the mass of a sample of a compound and the mass of each element in the sample.

÷ the mass of each element in the sample by the mass of the sample. × each by 100.

In your textbook, read about empirical and molecular formulas.

Circle the letter of the choice that best answers the question.

3. Which information about a compound can you use to begin to determine the empirical and molecular formulas of the compound?

a. mass of the compound

b. number of elements in the compound

c. percent composition of the compound

d. volume of the compound

4. You have determined that a compound is composed of 0.300 moles of carbon and 0.600 moles of oxygen. What must you do to determine the mole ratio of the elements in the empirical formula of the compound?

a. Multiply each mole value by 0.300 mol.

b. Multiply each mole value by 0.600 mol.

c. Divide each mole value by 0.300 mol.

d. Divide each mole value by 0.600 mol.

5. The mole ratio of carbon to hydrogen to oxygen in a compound is 1 mol C : 2 mol H : 1 mol O. What is the empirical formula of the compound?

a. CHO

b. CH<sub>2</sub>O

c. C<sub>2</sub>H<sub>2</sub>O<sub>2</sub>

d. C<sub>2</sub>H<sub>2</sub>O<sub>3</sub>

6. You calculate the mole ratio of oxygen to aluminum in a compound to be 1.5 mol O : 1 mol Al. What should you do to determine the mole ratio in the empirical formula of the compound?

a. Multiply each mole value by 1.5.

b. Multiply each mole value by 2.

c. Divide each mole value by 1.5.

d. Divide each mole value by 2.

7. What is the relationship between the molecular formula and the empirical formula of a compound?

a. (molecular formula)/(empirical formula) =  $n$

b. molecular formula =  $\frac{\text{empirical formula}}{n}$

c. molecular formula = (empirical formula) $n$

d. molecular formula =  $\frac{n}{\text{empirical formula}}$

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**ML Morrison**



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